Test Review № 7

Conservation of Mass. Matter is neither created nor destroyed. During a chemical reaction the mass does not change. A properly written equation shows conservation of mass. For example, $AgNO_3(aq) + NaCl(aq) \rightarrow NaNO_3(aq) + AgCl(s)$.

SILVER NITRATE	SODIUM CHLORIDE	SODIUM NITRATE	SILVER CHLORIDE
AgNO₃	**NaCl**	**NaNO₃**	**AgCl**
Ag = 1 × 108 = 108	Na = 1 × 23 = 23	Na = 1 × 23 = 23	Ag = 1 × 108 = 108
N = 1 × 14 = 14	Cl = 1 × 35 = 35	N = 1 × 14 = 14	Cl = 1 × 35 = 35
O = 3 × 16 = 48	58	O = 3 × 16 = 48	143
170		85	

$$AgNO_3(aq) \quad + \quad NaCl(aq) \rightarrow \quad NaNO_3(aq) \quad + \quad AgCl(s)$$

170	+	58	=	85	+	143
		228				228

Balancing Equations. The equation at the top of the box to the right does *not* show conservation of mass. Starting with two molecules of hydrogen, as shown in the equation at the bottom of the box by writing a **coefficient** 2 in front of the hydrogen and forming two molecules of water by writing a **coefficient** 2 in front of the water shows conservation. Coefficients are used to **balance** equations. Coefficients make the number of atoms of each type the same on the reactant and product side. As a result, coefficients make the mass the same on the reactant and product side of the equation. Balancing is done by counting the number and type of atoms on the reactant and product side of the equation and making them equal.

$$H_2 + O_2 \rightarrow H_2O$$
$$2 + 32 \neq 18$$

$$2H_2 + O_2 \rightarrow 2H_2O$$
$$2(2) + 32 = 2(18)$$
$$36 = 36$$

Moles. A mole is a formula mass expressed in grams. (1 mole = 1 gram formula mass). Atomic mass units are too small to measure on a laboratory balance, but grams are not. An atom of carbon has a mass of 12 amu and a molecule of glucose has a mass of 180 amu. Each mass represents one particle. Since the mass ratios in formula masses and gram formula masses are the same (12 amu:180 amu::12 g:180 g), the ratio of particles must

Substance	Formula Mass	Gram Formula Mass
carbon	12 amu	12 g
sodium chloride (NaCl)	58 amu	58 g
glucose ($C_6H_{12}O_6$)	180 amu	180 g

1. $GFM = \dfrac{g}{mole}$
2. $g = GFM \times mole$
3. $mole = \dfrac{g}{GFM}$

still be the same (1mole:1 mole). The gram formula mass (GFM) is the number of grams in 1 mole. This results in the mathematical relationships shown above and to the right.

Stoichiometry. Stoichiometry is the branch of chemistry that deals with the application of the laws of definite proportions and of the conservation of mass and energy to chemical activity. It shows the quantitative relationship between constituents of a chemical reaction. Stoichiometric calculations are based on several assumptions. It is assumed that the reaction has no side reactions, the reaction goes to completion, and the reactants are completely consumed. One type of problems that can be solved stoichiometrically is based on the mole ratios of a balanced equation. A sample problem is shown to the right.

Sample Problem

How many moles of oxygen are consumed when 0.6 moles of hydrogen burns to produce water?

Step 1: Write a balanced equation and determine the mole ratios from the equation

Step 2: Identify the known and the unknown

	$2H_2(g)$	$+ O_2(g)$	$\rightarrow 2H_2O$
mole ratio	2	1	2
moles	known 0.6	unknown x	

Step 3: Set up a proportion and solve for the unknown

$$\Rightarrow \frac{2}{0.6 mol} = \frac{1}{x}$$
$$\Rightarrow 2x = 0.6 mol$$
$$\Rightarrow x = 0.3 mol$$

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