

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Modern Chemistry Chapter 4 Answers

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Modern Chemistry Raymond E. Davis, 1999 2000 2005 State Textbook Adoption Rowan Salisbury Introduction to Modern Chemistry Edward Florian Neuzil, 1968 *Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **National Geographic Answer Book** National Geographic, 2015 This far reaching reference is designed with many entry points and a visually engaging format to satisfy the curious browser the student researcher and the earnest knowledge seeker alike **Supplement for Modern Organic Chemistry** John D. Roberts, Marjorie C. Caserio, 1967

Instructor's Manual to Accompany Chemistry in the Modern World, Concepts and Applications Frank L. Wiseman, 1985

Organic Synthesis Michael Smith, 2016-11-22 *Organic Synthesis* Fourth Edition provides a reaction based approach to this important branch of organic chemistry Updated and accessible this eagerly awaited revision offers a comprehensive foundation for graduate students coming from disparate backgrounds and knowledge levels to provide them with critical working knowledge of basic reactions stereochemistry and conformational principles This reliable resource uniquely incorporates molecular modeling content problems and visualizations and includes reaction examples and homework problems drawn from the latest in the current literature In the Fourth Edition the organization of the book has been improved to better serve students and professors and accommodate important updates in the field The first chapter reviews basic retrosynthesis conformations and stereochemistry The next three chapters provide an introduction to and a review of functional group exchange reactions these are followed by chapters reviewing protecting groups oxidation and reduction reactions and reagents hydroboration selectivity in reactions A separate chapter discusses strategies of organic synthesis and the book then delves deeper in teaching the reactions required to actually complete a synthesis Carbon carbon bond formation reactions using both nucleophilic carbon reactions are presented and then electrophilic carbon reactions followed by pericyclic reactions and radical and carbene reactions The important organometallic reactions have been consolidated into a single chapter Finally the chapter on combinatorial chemistry has been removed from the strategies chapter and placed in a separate chapter along with valuable and forward looking content on green organic chemistry process chemistry and continuous flow chemistry Throughout the text *Organic Synthesis* Fourth Edition utilizes Spartan generated molecular models class tested content and useful pedagogical features to aid student study and retention including Chapter Review Questions and Homework Problems A full Solutions Manual is also available online for qualified instructors to support teaching Winner 2018 Textbook Excellence Award Texty from the Textbook and Academic Authors Association Fully revised and updated throughout and organized into 19 chapters for a more cogent and versatile presentation of concepts Includes reaction examples taken from literature research reported between 2010 2015 Features new full color art and new chapter content on process chemistry and green organic chemistry Offers valuable study and teaching tools including Chapter Review Questions and Homework Problems for students Solutions Manual for qualified course instructors *Student's*

Guide to Chemistry; a Modern Introduction David Brooks,1974 **The Chautauquan** ,1889 **Operation of Wastewater Treatment Plants** ,2004 **Modern University Chemistry** Norbert T. Porile,1987 Modern Spectroscopy J. Michael Hollas,2013-04-11 The latest edition of this highly acclaimed title introduces the reader to a wide range of spectroscopies and includes both the background theory and applications to structure determination and chemical analysis It covers rotational vibrational electronic photoelectron and Auger spectroscopy as well as EXAFs and the theory of lasers and laser spectroscopy A revised and updated edition of a successful clearly written book Includes the latest developments in modern laser techniques such as cavity ring down spectroscopy and femtosecond lasers Provides numerous worked examples calculations and questions at the end of chapters **A Guide to Modern Chemistry** John F. Baxter,Jay A. Young,1959

Customized Modern Inorganic Chemistry William L. Jolly,1998 **The American Magazine** ,1913 American Magazine ,1913 **Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry** Iris Stovall,Donald J. DeCoste,2000 **Introductory Chemistry** Steve Russo, Mike Silver,2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support onThe Chemistry Place tm Special Edition and onThe Chemistry of Life CD ROM For college instructors and students

American Illustrated Magazine ,1913 *Chemistry: Principles and Applications* Curtis B. Anderson,Peter C. Ford,John H. Kennedy,John Harvey Kennedy,1973

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