



### **Friends Can Show Kindness**

**Friends can show kindness to each other in many ways.**

**You can show kindness by taking turns when you play.**

**Ask your friend if they would like to go "first".**

**If a friend gets hurt, you can help them get up.**

**You may need to go get a grownup to come and help.**

**When a friend is feeling sad, you can cheer them up.**

**Sometimes a friend just needs someone to listen.**

**You can invite a lonely child to come and play with you.**

**If you hear kids making fun of someone, say "Stop!"**

**Think of how you would feel if it happened to you.**

**Showing kindness to others is an important thing to do.**

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# Friendship Stories Kids Training Guide

**Patrick Vollmar**



## Friendship Stories Kids Training Guide:

**Resources in Education** ,1998-05    **Manual of Moral and Humane Education** Flora Helm Krause,1910  
*Teacher's Manual for Work Books in Citizenship Training for Children ... of the Elementary School* Walter Lansing  
Collins,1928    *A Second Manual for Training in Worship* Hugh Hartshorne,1921    **Training for World Friendship** Ina  
Corinne Brown,1929    **Proceedings of the Friends' General Conference** Society of Friends. Friends General  
Conference,1916    *A Guide to Good Reading* Robert Newton Linscott,1912    **Friends' Intelligencer** ,1925  
**Temperance Advocate and Seamen's Friend** ,1917 Vol 12 n s v 4 no 12 Dec 1855 has bound after it The Folio Nov 16  
1855    **Willing's Press Guide and Advertisers' Directory and Handbook** ,1904    **The Child Welfare Manual** ,1915  
**The American Friend** ,1930    The Booklist ,1928    **International Journal of Religious Education** ,1924  
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Friend* ,1881    *Farmers' Guide* ,1897    **A Kid's Guide to Making Friends** Joy Wilt Berry,1979 A children s book about  
social skills    **Ontario Library Review and Book-selection Guide** ,1924

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable

measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the  
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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(l) for which ΔH = -57 kJ mol<sup>-1</sup> eg 2 HNO<sub>3</sub>(aq) + KOH(aq) → KNO<sub>3</sub>(aq) + H<sub>2</sub>O(l) for which ΔH = -57 kJ mol<sup>-1</sup>  
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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(aq) you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts

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