

Preparation of a standard solution

Aim: To prepare 2.0 mol.dm^{-3} standard sodium chloride solution using a 250 cm^3 volumetric flask.

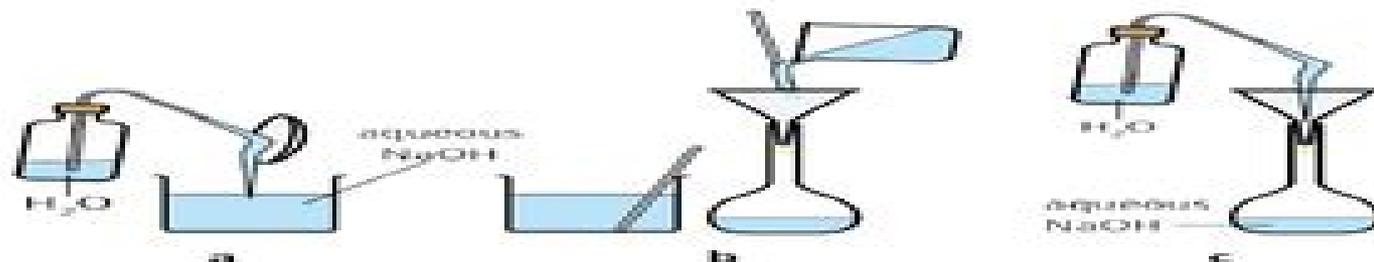
Tools 1: Experimental techniques – Applying techniques.

Available Apparatus: digital balance, volumetric flask, pipettes, and pipette bulbs, spatula, watch glass, beaker, funnel, stirring rod, and wash bottle.

Reagents: Sodium chloride, food coloring.

Preparation of sodium hydroxide solution.

1. Calculate the mass (in grams) of solid sodium hydroxide required to prepare 2.0 mol.dm^{-3} solution in the assigned volumetric flask.
2. Place a watch glass on the analytical balance and press the 'Tare' button.
3. Accurately weigh approximately the amount of sodium chloride, you have calculated you need, recording the exact amount used.
4. Add approximately 50 cm^3 distilled water to the beaker and stir with a glass rod until the solid has dissolved.
5. Using a funnel, carefully transfer your solution to the volumetric flask given.
6. Add approximately 10 cm^3 distilled water to the beaker, rinse it around the beaker and add to the volumetric flask. Repeat with a further 10 cm^3 portion.
7. Carefully add distilled water to the volumetric flask, running it around the edge of the funnel to wash any remaining NaCl in, and also running it over the stirring rod to wash this in.
8. When the level of solution is within 1 cm of the graduation on the flask, add distilled water drop-wise using a pipette until the bottom of the meniscus is in line with the graduation mark.
9. Firmly place the lid on the volumetric flask, and invert it several times, holding the lid in place. You will know your solution is fully mixed when inverting it no longer makes it look 'swirly'.
10. Calculate the exact concentration (4 s.f.) of the solution, using the mass measured in Step 2 and label the flask with the solution, the concentration, the date of preparation and your name.
11. The solution should be stored in a cool dark place until it is to be used.



A. Calculate the mass of sodium chloride required to prepare a solution 2.0 mol.dm^{-3} solution.

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Preparation Of Solutions Lab Report

Clifford Lane



Preparation Of Solutions Lab Report:

Whispering the Secrets of Language: An Psychological Journey through **Preparation Of Solutions Lab Report**

In a digitally-driven world where monitors reign supreme and immediate interaction drowns out the subtleties of language, the profound strategies and mental subtleties concealed within words usually move unheard. However, nestled within the pages of **Preparation Of Solutions Lab Report** a interesting fictional treasure sporting with organic emotions, lies an extraordinary quest waiting to be undertaken. Penned by an experienced wordsmith, that charming opus invites viewers on an introspective journey, delicately unraveling the veiled truths and profound impact resonating within ab muscles cloth of every word. Within the psychological depths of this emotional evaluation, we shall embark upon a genuine exploration of the book is core styles, dissect its fascinating publishing style, and yield to the powerful resonance it evokes heavy within the recesses of readers hearts.

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