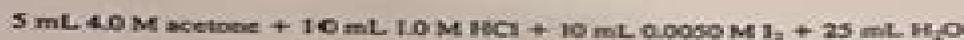


Experiment 20

Advance Study Assignment: The Iodination of Acetone

1. In a reaction involving the iodination of acetone, the following volumes were used to make up the reaction mixture:



- a. How many moles of acetone were in the reaction mixture? Recall that, for a compound A, moles $A = M_A \times V$, where M_A is the molarity of A and V is the volume in liters of the solution of A that was used. $5 \text{ mL} = 0.005 \text{ L}$

$$A = 4.0 \text{ M} \times 0.005 \text{ L}$$

$$\underline{0.02} \text{ moles acetone}$$

$$A = 0.02 \text{ moles acetone}$$

- b. What was the molarity of acetone in the reaction mixture? The volume of the mixture was 50 mL, 0.050 L, and the number of moles of acetone was found in Part (a). Again,

$$M_A = \frac{\text{moles of A}}{V \text{ of soln. in liters}}$$

$$M_{\text{H}} = \frac{0.02 \text{ M}}{0.050 \text{ L}} = 0.4 \text{ M acetone}$$

$$\underline{0.4} \text{ M acetone}$$

- c. How could you double the molarity of the acetone in the reaction mixture, keeping the total volume at 50 mL and keeping the same concentrations of H^+ ion and I_2 as in the original mixture?

1. Double the moles of acetone concentration

2. Using the reaction mixture in Problem 1, a student found that it took 510 seconds for the color of the I_2 to disappear.

- a. What was the rate of the reaction? Hint: First find the initial concentration of I_2 in the reaction mixture, $[I_2]_0$. Then use Equation 3.

$$[I_2]_0 = 0.0050 \text{ M} \times 0.01 = \frac{0.000050 \text{ mol}}{0.050 \text{ L}} = 0.001 \text{ M}$$

$$\text{rate} = \underline{2.0 \times 10^{-6}} \text{ M/sec}$$

$$\text{rate} = \frac{0.001 \text{ M}}{510 \text{ s}} = 1.96 \times 10^{-6} \text{ M} \cdot \text{s}^{-1} \approx 2.0 \times 10^{-6} \text{ M} \cdot \text{s}^{-1}$$

- b. Given the rate from Part (a), and the initial concentrations of acetone, H^+ ion, and I_2 in the reaction mixture, write Equation 3 as it would apply to the mixture.

$$\text{rate} = k [\text{acetone}]^m [I_2]^n [H^+]^p$$

$$2.0 \times 10^{-6} = k (0.4 \text{ M})^m (0.02)^n (0.001)^p$$

- c. What are the unknowns that remain in the equation in Part (b)?

$$\underline{k \quad m \quad n \quad p}$$

(continued on following page)

$$\begin{aligned} \text{mol } H^+ &= 1.0 \text{ M} \times 0.010 \\ &= \frac{0.010}{0.050} = 0.20 \end{aligned}$$

Advance Study Assignment Experiment 30 Answers

SJ Ball



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Mar 15 1990 thunder cake is a danish contemporary realistic fiction book that tells the story of a young girl and her grandmother trying to forget the powerful thunderstorm coming in on them by making a cake this book would be appealing to children as for some of them are quite scared of thunderstorms

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a loud clap of thunder booms and rattles the windows of grandma s old farmhouse this is thunder cake baking weather calls grandma as she and her granddaughter hurry to gather the ingredients around the farm a real thunder cake

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