

Acid-Base Titration Practice Problems

- 1) Calculate the pH during the titration of 40.00 mL of 0.0500 M butanoic acid, $\text{CH}_3\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$, with 0.1000 M NaOH solution after the following additions of the base: a) 0.00 mL; b) 15.00 mL; c) 19.00 mL; d) 20.00 mL; e) 25.00 mL. The K_a for butanoic acid is 1.54×10^{-5} .
- 2) Calculate the pH during the titration of 25.00 mL of 0.200 M KOH with 0.100 M HClO_4 after the following additions of acid: a) 0.00 mL; b) 25.00 mL; c) 39.00 mL; d) 50.00 mL; e) 60.00 mL.
- 3) Calculate the pH when the following quantities of 0.0500 M $\text{Ba}(\text{OH})_2$ solution have been added to 25.00 mL of 0.0400 M HNO_3 solution: a) 0.00 mL; b) 5.00 mL; c) 10.00 mL; d) 20.00 mL.

Acid Base Titration Problems With Answers

Lei Huang



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web solution a first order with respect to br first order with respect to bro₃ and second order with respect to h hence the overall order of the reaction is equal to 1 + 1 + 2 = 4 b order of the reaction with respect to acetaldehyde is 3 + 2 and overall order is also 3 + 2

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web in order to have successful collisions the colliding particles must have both the proper amount of energy and the proper fill in blank draw an energy diagram which shows the shape of the curve you might expect for the reaction in this question the overall reaction is exothermic

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web mar 13 2023 work in groups on these problems you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 for the reaction below if substance a is disappearing at a rate of 1.82 times 10^{-2} mol l⁻¹ s⁻¹ at what rate is c appearing 3 a 3 b \rightarrow 5 c 2 d

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web reaction rates google classroom mg s 2 hcl a q mgcl a 2 a q h a 2 g the reaction between mg s and hcl a q is represented by the equation above in a kinetics experiment a 0.080 g sample of mg s ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 m hcl a q at 25 c

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