

## Acid-Base Titration Practice Problems

- 1) Calculate the pH during the titration of 40.00 mL of 0.0500 M butanoic acid,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$ , with 0.1000 M NaOH solution after the following additions of the base: a) 0.00 mL; b) 15.00 mL; c) 19.00 mL; d) 20.00 mL; e) 25.00 mL. The  $K_a$  for butanoic acid is  $1.54 \times 10^{-5}$ .
- 2) Calculate the pH during the titration of 25.00 mL of 0.200 M KOH with 0.100 M  $\text{HClO}_4$  after the following additions of acid: a) 0.00 mL; b) 25.00 mL; c) 39.00 mL; d) 50.00 mL; e) 60.00 mL.
- 3) Calculate the pH when the following quantities of 0.0500 M  $\text{Ba}(\text{OH})_2$  solution have been added to 25.00 mL of 0.0400 M  $\text{HNO}_3$  solution: a) 0.00 mL; b) 5.00 mL; c) 10.00 mL; d) 20.00 mL.

# Acid Base Titration Practice Problems With Answers

**Scott C. Dulebohn**



**Acid Base Titration Practice Problems With Answers:**

## Unveiling the Magic of Words: A Report on "**Acid Base Titration Practice Problems With Answers**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Acid Base Titration Practice Problems With Answers**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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