

## Titration Calculation Practice Questions

1. Eggshells contain calcium carbonate. The following procedure was used to determine the percentage, by mass, of calcium carbonate in 1.13 g of eggshell.

**Step 1** The 1.13 g of eggshell were crushed and added to 100.0 mL of 0.300 mol L<sup>-1</sup> HCl solution. Excess HCl remained after the reaction was complete.



**Step 2** 20.0 mL samples of the excess HCl were titrated with 0.10 mol L<sup>-1</sup> Na<sub>2</sub>CO<sub>3</sub> solution.

- (a) Calculate the initial number of moles of HCl present before the reaction in Step 1.  
(b) The equation for the titration reaction in Step 2 is shown below:



In one titration 8.35 mL of Na<sub>2</sub>CO<sub>3</sub> was needed to neutralise the HCl in one 20.0 mL sample.

- (i) Calculate the number of moles of Na<sub>2</sub>CO<sub>3</sub> needed to neutralise the HCl in Step 2.  
(ii) Hence state the number of moles of HCl in the 20.0 mL sample titrated in Step 2.  
(iii) Hence calculate the total number of moles of excess HCl that remained after the reaction with the eggshell in Step 1.  
(c) Calculate the number of moles of HCl that reacted with the eggshell in Step 1.  
(d) Calculate the mass of calcium carbonate in the 1.13 g of eggshell.  
(e) Calculate the percentage mass of calcium carbonate in eggshell.

2. The following procedure was used to determine the concentration of hydrogen gas in a sample of air:

**Step 1** 1.0 × 10<sup>2</sup> L of air was bubbled through 0.0500 L of 0.300 mol L<sup>-1</sup> acidified K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> solution. An equation for the reaction that occurred is shown below:



Excess Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> remained in the solution after the reaction.

**Step 2** The excess Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> was titrated with 2.00 mol L<sup>-1</sup> Fe<sup>2+</sup> solution. An equation for the reaction that occurred is shown below:



A titre value of 17.75 mL was obtained.

- (a) Calculate the number of moles of Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> present before the reaction with H<sub>2</sub> in Step 1.  
(b) Calculate the number of moles of Fe<sup>2+</sup> required to react with the Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> in Step 2.  
(c) Hence calculate the number of moles of Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> left unreacted after Step 1.  
(d) Hence calculate the number of moles of Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> that reacted with H<sub>2</sub> in Step 1.  
(e) Calculate the number of moles of H<sub>2</sub> in the 1.0 × 10<sup>2</sup> L of air.  
(f) Calculate the concentration, in %w/v, of hydrogen gas in the air.

# Titration Questions And Answers

**Lei Huang**



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## **Table of Contents Titration Questions And Answers**

1. Understanding the eBook Titration Questions And Answers
  - The Rise of Digital Reading Titration Questions And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Titration Questions And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Titration Questions And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Titration Questions And Answers
  - Personalized Recommendations
  - Titration Questions And Answers User Reviews and Ratings
  - Titration Questions And Answers and Bestseller Lists
5. Accessing Titration Questions And Answers Free and Paid eBooks

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  - Titration Questions And Answers eBook Subscription Services
  - Titration Questions And Answers Budget-Friendly Options
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    - ePub, PDF, MOBI, and More
    - Titration Questions And Answers Compatibility with Devices
    - Titration Questions And Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Titration Questions And Answers
    - Highlighting and Note-Taking Titration Questions And Answers
    - Interactive Elements Titration Questions And Answers
  8. Staying Engaged with Titration Questions And Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Titration Questions And Answers
  9. Balancing eBooks and Physical Books Titration Questions And Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Titration Questions And Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Titration Questions And Answers
    - Setting Reading Goals Titration Questions And Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Titration Questions And Answers
    - Fact-Checking eBook Content of Titration Questions And Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web dans le tome 2 les fêlures du mâles il nage des bas fonds anversois aux salons de la périphérie sans se noyer dans l hypocrisie ses amies lui dessillent les yeux lui ouvrent l esprit et la braguette si le mot cul vous dérange dans une phrase quelle que soit la beauté du texte tournez la page et vous aurez avalé le bouquin d un