

Chemistry Notes 14.3

Three Basic Assumptions of the kinetic theory:

- 1) Gas is composed of particles, usually molecules or atoms
- 2) Particles in a gas move rapidly in constant random motion (they move in straight line paths only colliding with one another or other objects)
- 3) Collisions are perfectly elastic: kinetic energy is transferred without loss from one particle to another

Equation: $P \times V = n \times R \times T$

$R = 8.31 \text{ (L} \times \text{kPa) / (mol} \times \text{K)}$

$n = (P \times V) / (R \times T)$

Ideal Gas:

- 1) A gas is composed of particles with a mass that is low enough to be considered almost zero.
 - But a real gas has limitations to the pressure it can be compressed to before it condenses into a liquid

Real gases differ from ideal gases in the following ways:

- Real gases have small attractive and repulsive forces between particles, and ideal gases do not
- Real gas particles have a volume, and ideal gas particles do not
- Real gas particles collide inelastically and lose energy with collisions, and ideal gas particles collide elastically
- Real gases are not infinitely compressible

Ideal gases don't exist because they do take up space, and there are attractive forces

Diffusion: molecules moving from areas of high concentration to low concentration

- Perfume

Effusion: Gas escaping through tiny hole in container

- Both of these depend on the molar mass of the particle, which determines the speed
- The rate of effusion and diffusion is inversely proportional to the square root of the molar mass

To solve Graham's law problems, put whatever you're looking for in the numerator
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