

# CHEMISTRY

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## PERIODIC TABLE OF THE ELEMENTS



### ATOMIC STRUCTURE

Atomic Number,  $Z$ , of an element is the number of protons in the nucleus. For a neutral atom,  $Z = +$  the number of electrons. The mass number,  $A$ , is the sum of the number of protons and neutrons in the nucleus. Isotopes of an element have the same  $Z$ , different  $A$ . The atomic mass is the weighted average of isotopes. See the periodic table for a number of isotopes. In practical work use the atomic mass number or atomic weight (weighted average of isotopes) for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number,  $n$ , is the number of shells.  $n = 1, 2, 3, \dots$ . The number of orbitals in a shell is  $n^2$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ .

### NUCLEAR CHEMISTRY

Nuclear reactions alter the number, protons, neutrons, or mass of a nucleus. Nuclear reactions are used to produce energy. Nuclear reactions are used to produce energy. Nuclear reactions are used to produce energy.

### MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in atoms. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy.

n	l	m <sub>l</sub>	m <sub>s</sub>	Orbitals
1	0	0	±1/2	1 orbital
2	0	0	±1/2	1 orbital
2	1	0, ±1	±1/2	3 orbitals
3	0	0	±1/2	1 orbital
3	1	0, ±1	±1/2	3 orbitals
3	2	0, ±1, ±2	±1/2	5 orbitals
4	0	0	±1/2	1 orbital
4	1	0, ±1	±1/2	3 orbitals
4	2	0, ±1, ±2	±1/2	5 orbitals
4	3	0, ±1, ±2, ±3	±1/2	7 orbitals

Atomic configuration: How orbitals are filled. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. Aufbau Principle: Electrons fill orbitals in order of increasing energy.



# Chemistry Quickstudy Reference Guides Academic

**Daniela Niemeyer**



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## **Table of Contents Chemistry Quickstudy Reference Guides Academic**

1. Understanding the eBook Chemistry Quickstudy Reference Guides Academic
  - The Rise of Digital Reading Chemistry Quickstudy Reference Guides Academic
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Quickstudy Reference Guides Academic
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Quickstudy Reference Guides Academic
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Quickstudy Reference Guides Academic
  - Personalized Recommendations
  - Chemistry Quickstudy Reference Guides Academic User Reviews and Ratings
  - Chemistry Quickstudy Reference Guides Academic and Bestseller Lists
5. Accessing Chemistry Quickstudy Reference Guides Academic Free and Paid eBooks
  - Chemistry Quickstudy Reference Guides Academic Public Domain eBooks
  - Chemistry Quickstudy Reference Guides Academic eBook Subscription Services

- Chemistry Quickstudy Reference Guides Academic Budget-Friendly Options
- 6. Navigating Chemistry Quickstudy Reference Guides Academic eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Quickstudy Reference Guides Academic Compatibility with Devices
  - Chemistry Quickstudy Reference Guides Academic Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Quickstudy Reference Guides Academic
  - Highlighting and Note-Taking Chemistry Quickstudy Reference Guides Academic
  - Interactive Elements Chemistry Quickstudy Reference Guides Academic
- 8. Staying Engaged with Chemistry Quickstudy Reference Guides Academic
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Quickstudy Reference Guides Academic
- 9. Balancing eBooks and Physical Books Chemistry Quickstudy Reference Guides Academic
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Quickstudy Reference Guides Academic
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Quickstudy Reference Guides Academic
  - Setting Reading Goals Chemistry Quickstudy Reference Guides Academic
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Quickstudy Reference Guides Academic
  - Fact-Checking eBook Content of Chemistry Quickstudy Reference Guides Academic
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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