

## Chemistry 338 2<sup>nd</sup> Semester Final Exam Review Sheet

### Unit 1 - Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100$$

$$\% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

### Unit 2 – Periodic Table – Chapter 5

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten ion	$^{184}_{74}\text{W}^{+2}$	74	74	110	72	184	+2

3. Using the following atomic radius values, predict the atomic radius of potassium (K):  
Li – 145 pm, Na – 180 pm, K - ?, Rb – 235 pm  
K – 220 pm

### Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
- $\text{Fe}_2\text{O}_3$  - iron (III) oxide (ionic - roman numerals for transition metals, change ending to -ide)
  - $\text{Cl}_2\text{O}_8$  - dichlorine octaoxide (covalent - use prefixes, change ending to -ide)
5. Classify the compound as ionic or covalent. Write the formula
- Trisulfur Dinitride –  $\text{S}_3\text{N}_2$  (covalent)
  - Zinc Hydroxide –  $\text{Zn}(\text{OH})_2$  (ionic - look up ion charges and criss-cross)  $\text{Zn}^{+2} \text{OH}^{-1}$
  - Chromium (III) nitrate –  $\text{Cr}(\text{NO}_3)_3$  (ionic - criss-cross)  $\text{Cr}^{+3} \text{NO}_3^{-2}$
6. Write the equation, balance it, and classify the type of reaction:  
Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
- $$1 \text{ FeSO}_4 + 2 \text{ NaOH} \rightarrow 1 \text{ Fe}(\text{OH})_2 + 1 \text{ Na}_2\text{SO}_4$$
- double displacement (two ionic compounds reacting)
7. Balance the equation and classify the type of reaction:
- $$1 \text{ Cu} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Cu}(\text{NO}_3)_2 + 2 \text{ Ag}$$
- Single displacement (element and a compound reacting)

### Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of  $\text{Fe}_2(\text{SO}_4)_3$ .
- $$\begin{aligned} \text{Fe } (2)(55.9 \text{ g/mol}) &= 111.8 \text{ g/mol} \\ \text{S } (3)(32.1 \text{ g/mol}) &= 96.3 \text{ g/mol} \\ + \text{O } (12)(16.0 \text{ g/mol}) &= +192 \text{ g/mol} \\ &= 400.1 \text{ g/mol} \end{aligned}$$

# Chemistry Spring Semester Final Exam Review Sheet

**Jicheng Xie**



**Chemistry Spring Semester Final Exam Review Sheet:**

## Decoding **Chemistry Spring Semester Final Exam Review Sheet**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Chemistry Spring Semester Final Exam Review Sheet**," a mesmerizing literary creation penned by a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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