

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- | | |
|-----------------------------|---|
| a. Sulfur hexafluoride | SF_6 |
| b. Lithium nitride | Li_3N |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$ |
| d. Tin (II) chloride | SnCl_2 |
| e. Ammonium acetate | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$) |
| f. Mercury (I) chloride | Hg_2Cl_2 |
| g. Potassium bromate | KBrO_3 |
| h. Hydrosulfuric acid | $\text{H}_2\text{S(aq)}$ |
| i. Chloric acid | HClO_3 |
| j. Sulfurous acid | H_2SO_3 |

Name each of the following compounds:

- | | |
|------------------------------|-------------------------|
| a. CuSO_4 | Copper (II) sulfate |
| b. AlF_3 | aluminum fluoride |
| c. HI | hydrogen iodide |
| d. NO | nitrogen monoxide |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid |
| f. HNO_3 | nitric acid |
| g. HNO_2 | nitrous acid |
| h. NaHSO_4 | sodium hydrogen sulfate |

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \rightarrow 20 \rightarrow 10 \rightarrow 5 \rightarrow 2.5 \rightarrow 1.25$$

1 2 3 4 5

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \rightarrow 50 \rightarrow 25 \rightarrow 12.5 \text{ mg}$$

1 2 3

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 44 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

Chemistry Chapter 8 Study Guide Answers

Richard Bailey



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Table of Contents Chemistry Chapter 8 Study Guide Answers

1. Understanding the eBook Chemistry Chapter 8 Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 8 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 8 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 8 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 8 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chapter 8 Study Guide Answers User Reviews and Ratings
 - Chemistry Chapter 8 Study Guide Answers and Bestseller Lists

5. Accessing Chemistry Chapter 8 Study Guide Answers Free and Paid eBooks
 - Chemistry Chapter 8 Study Guide Answers Public Domain eBooks
 - Chemistry Chapter 8 Study Guide Answers eBook Subscription Services
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6. Navigating Chemistry Chapter 8 Study Guide Answers eBook Formats
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 - Chemistry Chapter 8 Study Guide Answers Compatibility with Devices
 - Chemistry Chapter 8 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 8 Study Guide Answers
 - Highlighting and Note-Taking Chemistry Chapter 8 Study Guide Answers
 - Interactive Elements Chemistry Chapter 8 Study Guide Answers
8. Staying Engaged with Chemistry Chapter 8 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 8 Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 8 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 8 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 8 Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 8 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 8 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 8 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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