

CHAPTER 8 REVIEW

Chemical Equations and Reactions

SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

1. Match the symbol on the left with its appropriate description on the right.

 d

(a) A precipitate forms.

 a ↓

(b) A gas forms.

 b

(c) A reversible reaction occurs.

 f (f)

(d) Heat is applied to the reactants.

 e (aq)

(e) A chemical is dissolved in water.

 c

(f) A chemical is in the liquid state.

2. Finish balancing the equation:

following



3. In each of the following formulas, write the total number of atoms present.

12 atoms _____ a. 4SO_2

16 atoms _____ b. 3O_2

51 atoms _____ c. $3\text{Al}_2(\text{SO}_4)_3$

3×10^{24} atoms _____ d. $6 \times 10^2 \text{ } 3 \text{ HNO}_3$

4. Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride → aluminum fluoride + copper metal



5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.



6. a. Balance the following equation:



Chapter 8 Section 1 Answers

Rachel Sandford



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