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web google classroom for the following redox equation $\text{MnO}_4^- + \text{A} + \text{A} + \text{Q} + \text{SO}_2 + \text{G} + \text{Mn}^{2+} + \text{A} + \text{Q} + \text{HSO}_4^- + \text{A} + \text{A} + \text{Q}$ what will be the stoichiometric coefficient of HSO_4^- if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}^+ + \text{H}_2\text{O}_2 + 2\text{Au}^{3+}$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 + \text{H}_2\text{O}_2 + 2\text{OH}^- + 3\text{Au}^{3+}$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^{-}$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^{-} + 14\text{H}^{+} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^{-}$ equalize the electrons $6\text{e}^{-} + 14\text{H}^{+} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^{-}$ multiplied by a factor of 6

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