

11

CHEMICAL REACTIONS

Chapter Test B

A. Matching

Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line.

Column A	Column B
_____ 1. a reaction in which an element or compound reacts with oxygen, often producing energy in the form of heat or light	a. decomposition reaction
_____ 2. a reaction in which two or more substances react to form a single substance	b. activity series
_____ 3. an equation that indicates only those particles that actually take part in the reaction	c. spectator ions
_____ 4. a substance that speeds up a reaction without being used up	d. balanced equation
_____ 5. ions that are not directly involved in a reaction	e. double-replacement reaction
_____ 6. a reaction in which a single compound is broken down into two or more products	f. catalyst
_____ 7. an equation in which each side has the same number of atoms of each element	g. combustion reaction
_____ 8. a reaction in which atoms of an element replace the atoms of a second element in a compound	h. combination reaction
_____ 9. a list of metals in order of decreasing reactivity	i. net ionic equation
_____ 10. a reaction that involves an exchange of positive ions between two compounds	j. single-replacement reaction

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. In the chemical equation $2\text{H}_2\text{O}_2(\text{aq}) \xrightarrow{\text{MnO}_2} 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$, the MnO_2 is a:
- a. reactant.
b. product.
c. spectator ion.
d. catalyst.
- _____ 12. When the equation $\text{Mg}(\text{s}) + \text{HCl}(\text{aq}) \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2(\text{g})$ is balanced, what is the coefficient for HCl?
- a. 1
b. 2
c. 3
d. 4

Pearson Chapter 11 Test

Julia Schneider



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Table of Contents Pearson Chapter 11 Test

1. Understanding the eBook Pearson Chapter 11 Test
 - The Rise of Digital Reading Pearson Chapter 11 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chapter 11 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chapter 11 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chapter 11 Test
 - Personalized Recommendations
 - Pearson Chapter 11 Test User Reviews and Ratings
 - Pearson Chapter 11 Test and Bestseller Lists
5. Accessing Pearson Chapter 11 Test Free and Paid eBooks
 - Pearson Chapter 11 Test Public Domain eBooks
 - Pearson Chapter 11 Test eBook Subscription Services
 - Pearson Chapter 11 Test Budget-Friendly Options
6. Navigating Pearson Chapter 11 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chapter 11 Test Compatibility with Devices
 - Pearson Chapter 11 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chapter 11 Test
 - Highlighting and Note-Taking Pearson Chapter 11 Test
 - Interactive Elements Pearson Chapter 11 Test
8. Staying Engaged with Pearson Chapter 11 Test

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Pearson Chapter 11 Test
- 9. Balancing eBooks and Physical Books Pearson Chapter 11 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chapter 11 Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chapter 11 Test
 - Setting Reading Goals Pearson Chapter 11 Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chapter 11 Test
 - Fact-Checking eBook Content of Pearson Chapter 11 Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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