

General Chemistry Questions

Electronic Structure and Periodic Table

1. What value or values of m_l are allowable for an orbital with $l = 2$?
 - a. 0
 - b. 2
 - c. -1
 - d. none of the above
 - e. all of the above
2. According to Bohr Theory, which of the following transitions in the hydrogen atom will give rise to the **least energetic** photon?
Use the equation: $E_n = (-2.18 \times 10^{-18} \text{ J})(1/n^2)$
 - a. $n = 5$ to $n = 3$
 - b. $n = 6$ to $n = 1$
 - c. $n = 4$ to $n = 3$
 - d. $n = 5$ to $n = 4$
 - e. $n = 6$ to $n = 5$
3. Consider a $3d_{zz}$ orbital. Which of the following statements is **incorrect**?
 - a. The xz plane is a nodal surface.
 - b. The xz plane divides the electron probability distribution into two identical mirror-image halves.
 - c. The xy plane divides the electron probability distribution into two identical mirror-image halves.
 - d. The yz plane divides the electron probability distribution into two identical mirror-image halves.
 - e. The nucleus is located at a node.
4. The electronic configuration of the element whose atomic number is 26 is:
 - a. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^0 3d^6$
 - b. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$
 - c. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
 - d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6 4p^2$
 - e. none of the above
5. Which of the following has the largest radius?
 - a. F
 - b. N
 - c. C
 - d. O
 - e. Ne

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Mary Burnham, Carol Hurd, 1928

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