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Thermodynamics

- Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
 - ΔS is negative and therefore, ΔH should be highly positive
 - ΔS is negative and therefore, ΔH should be highly negative
 - ΔS is positive and therefore, ΔH should be negative
 - ΔS is positive and therefore, ΔH should also be highly positive
- For the reaction, $X_2O_4(l) \longrightarrow 2XO_2(g)$, $\Delta U = 2.1 \text{ kcal}$, $\Delta S = 20 \text{ cal K}^{-1}$ at 300 K. Hence, ΔG is [2014]
 - 2.7 kcal
 - 2.7 kcal
 - 9.3 kcal
 - 9.3 kcal
- For a given exothermic reaction, K_p and K_p' are the equilibrium constants at temperatures T_1 and T_2 , respectively. Assuming that heat of reaction is constant in temperature range between T_1 and T_2 , it is readily observed that [2014]
 - $K_p > K_p'$
 - $K_p < K_p'$
 - $K_p = K_p'$
 - $K_p = \frac{1}{K_p'}$
- A reaction having equal energies of activation for forward and reverse reactions has [2013]
 - $\Delta S = 0$
 - $\Delta G = 0$
 - $\Delta H = 0$
 - $\Delta H = \Delta G = \Delta S = 0$
- In which of the following reactions, standard reaction entropy changes (ΔS°) is positive and standard Gibbs energy change (ΔG°) decreases sharply with increasing temperature? [2012]
 - $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$
 - $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
 - $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
 - $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
- The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
 - 10.52 cal/mol K
 - 21.04 cal/mol K
 - 5.260 cal/mol K
 - 0.526 cal/mol K
- Standard enthalpy of vaporisation $\Delta_{\text{vap}}H^\circ$ for water at 100°C is 40.66 kJ mol⁻¹. The internal energy of vaporisation of water at 100°C (in kJ mol⁻¹) is (assume water vapour to behave like an ideal gas). [2012]
 - + 37.56
 - 43.76
 - + 43.76
 - + 40.66
- If the enthalpy change for the transition of liquid water to steam is 30 kJ mol⁻¹ at 27°C, the entropy change for the process would be [2011]
 - 1.0 J mol⁻¹K⁻¹
 - 0.1 J mol⁻¹K⁻¹
 - 100 J mol⁻¹K⁻¹
 - 10 J mol⁻¹K⁻¹

Engineering Thermodynamics Important Questions Answers

Thomas A. Brzustowski



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Chemical Engineering Thermodynamics RAO, Y. V. C. Rao, 1997 *Fundamentals of Engineering Thermodynamics* R. Yadav, 2023-05-10 Thermodynamics deals with energy interactions between material bodies. It is the science of 3E's namely Energy, Entropy, and Equilibrium. The applications of its laws and principles are found in all fields of energy technology, notably in steam gas and nuclear power plants, internal combustion engines, gas turbines, jet propulsion, refrigeration, air conditioning, compressors, gas dynamics, and direct energy conversion. Starting with the basic concept, the book discusses the important topics such as basic concepts, heat and work, energy, ideal and real gases, zeroth, first, and second laws of thermodynamics, entropy, and third law, available energy and exergy, gas power cycles, vapour power cycles, general thermodynamic relations, refrigeration cycles, psychrometry, non-reactive mixtures, reactive mixture, chemical equilibrium, direct energy conversion, compressible flows, and heat transfer. The book is an essential text for BE/B.Tech for Mechanical Engineering students, UPSC, and GATE examinations.

Applied Chemical Engineering Thermodynamics Dimitrios P. Tassios, 2013-12-19 *Applied Chemical Engineering Thermodynamics* provides the undergraduate and graduate student of chemical engineering with the basic knowledge, the methodology, and the references he needs to apply it in industrial practice. Thus, in addition to the classical topics of the laws of thermodynamics, pure component and mixture thermodynamic properties, as well as phase and chemical equilibria, the reader will find history of thermodynamics, energy conservation, intermolecular forces, and molecular thermodynamics, cubic equations of state, statistical mechanics. A great number of calculated problems with solutions and an appendix with numerous tables of numbers of practical importance are extremely helpful for applied calculations. The computer programs on the included disk help the student to become familiar with the typical methods used in industry for volumetric and vapor-liquid equilibria calculations.

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use concerns all of us in general and the engineers of today and tomorrow in particular The study of thermodynamics the science of energy is a critical element in the education of all types of engineers Engineering Thermodynamics provides a thorough introduction to the art and science of engineering thermodynamics It describes in a straightforward fashion the basic tools necessary to obtain quantitative solutions to common engineering applications involving energy and its conversion conservation and transfer This book is directed toward sophomore junior and senior students who have studied elementary physics and calculus and who are majoring in mechanical engineering it serves as a convenient reference for other engineering disciplines as well The first part of the book is devoted to basic thermodynamic principles essentially presented in the classic way the second part applies these principles to many situations including air conditioning and the interpretation of statistical phenomena

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