

Chapter 17: Chemical Equilibrium

Worksheet: K_c and K_p

1. For the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$

at 1000 K

- 1. The equilibrium constant K_c for the reaction $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$ is 1.6×10^3 . Calculate K_c for the reaction $\text{SO}_2(g) + \frac{1}{2}\text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$.
- 2. For the reaction $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$, $K_c = 0.50$ at 400°C. Calculate K_c for the reaction $\frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g) \rightleftharpoons \text{NH}_3(g)$.
- 3. For the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$, $K_c = 4.0$ at 1500 K. Calculate K_c for the reaction $\text{NO}(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{1}{2}\text{O}_2(g)$.
- 4. For the reaction $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$, $K_c = 1.6$ at 1000 K. Calculate K_c for the reaction $2\text{CO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{CO}_2(g)$.
- 5. For the reaction $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$, $K_c = 50.0$ at 400°C. Calculate K_c for the reaction $\text{HI}(g) \rightleftharpoons \frac{1}{2}\text{H}_2(g) + \frac{1}{2}\text{I}_2(g)$.
- 6. For the reaction $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$, $K_c = 0.50$ at 400°C. Calculate K_c for the reaction $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$.
- 7. For the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$, $K_c = 4.0$ at 1500 K. Calculate K_c for the reaction $\text{NO}(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{1}{2}\text{O}_2(g)$.
- 8. For the reaction $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$, $K_c = 1.6$ at 1000 K. Calculate K_c for the reaction $2\text{CO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{CO}_2(g)$.
- 9. For the reaction $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$, $K_c = 50.0$ at 400°C. Calculate K_c for the reaction $\text{HI}(g) \rightleftharpoons \frac{1}{2}\text{H}_2(g) + \frac{1}{2}\text{I}_2(g)$.
- 10. For the reaction $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$, $K_c = 0.50$ at 400°C. Calculate K_c for the reaction $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$.

Answers

- 1. 1.6×10^3
- 2. 0.50
- 3. 0.25
- 4. 0.31
- 5. 0.050
- 6. 0.50
- 7. 0.25
- 8. 0.31
- 9. 0.050
- 10. 0.50

Elements Of Language First Course Answer Key

Marco Casella



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