

## 5.6 Writing Formulas for Ionic Compounds Assignment

### Instructions:

For this assignment, you are being asked to write formulas for Ionic Compounds and for Ionic Compounds containing Polyatomic Ions. Review the examples for each section and then write the formula for the identified compounds.

### Writing Formulas for Ionic Compounds Assignment

#### Examples:

Write the formula of potassium bromide.

- $K^{+1}$      $Br^{-1}$
- The sum of the charges is  $+1 + (-1) = 0$ , therefore, the formula is KBr

Write the formula of calcium bromide.

- $Ca^{+2}$      $Br^{-1}$
- The sum of the oxidation numbers is  $+2 + (-1) = +1$ , therefore, we need a second bromide (-1) to obtain a proper formula.
- The new sum is  $+2 + 2(-1) = 0$ , therefore, the formula is  $CaBr_2$ .

A simpler way to understand how to write formulas is to use the crisscross & reduce method. If there is only one atom of an element in the compound, you do not need to put the number 1 behind it. The following are examples:



#### Questions 1-6.

Write formulas for the following compounds:

1. sodium and fluorine



4. aluminum and sulfur



2. magnesium and iodine



5. hydrogen and oxygen



3. barium and oxygen



6. lithium and nitrogen



# Assignment 5 Ionic Compounds

**Lawrence Powell Eblin**



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