

# ARM Instruction Set Format

31	28:27	16:15	8:7	0					
Cond	0 0 1	Opcode	S	Rn	Rd	Operand2			
Cond	0 0 0 0 0 0	A S	Rd	Rn	Rs	1 0 0 1	Rm		
Cond	0 0 0 0 1	U A S	RdHi	RdLo	Rs	1 0 0 1	Rm		
Cond	0 0 0 1 0	B 0 0	Rn	Rd	0 0 0 0	1 0 0 1	Rm		
Cond	0 1 1	B U B W L	Rn	Rd	Offset				
Cond	1 0 0	B U S W L	Rn	Register List					
Cond	0 0 0	F U 1 W L	Rn	Rd	Offset1	1 S H 1	Offset2		
Cond	0 0 0	F U 0 W L	Rn	Rd	0 0 0 0	1 S H 1	Rm		
Cond	1 0 1 1	Offset							
Cond	0 0 0 1	0 0 1 0	1 1 1 1	1 1 1 1	1 1 1 1	0 0 0 1	Rn		
Cond	1 1 0	B U N W L	Rn	CRd	CPNum	Offset			
Cond	1 1 1 0	Op1	CRn	CRd	CPNum	Op2	0	CRm	
Cond	1 1 1 0	Op1	L	CRn	Rd	CPNum	Op2	1	CRm
Cond	1 1 1 1	SWI Number							

## Instruction type

Data processing / PSR Transfer

Multiply

Long Multiply (v3M / v4 only)

Swap

Load/Store Byte/Word

Load/Store Multiple

Halfword transfer: Immediate offset (v4 only)

Halfword transfer: Register offset (v4 only)

Branch

Branch Exchange (v4T only)

Coprocessor data transfer

Coprocessor data operation

Coprocessor register transfer

Software interrupt

# Arm Instruction Set Assembly Language Utk

**Roman Wölfel**



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subsequent  $\Delta h$  values  $2 \text{ CO}_2 \text{ (g)}$   $2 \text{ H}_2 \text{O (g)}$   $2 \text{ C}_2\text{H}_2 \text{ (g)}$   $5 \text{ O}_2 \text{ (g)}$   $2 \text{ H}_2 \text{ (g)}$   $2 \text{ C}_2\text{H}_6 \text{ (g)}$   $\Delta h = 94.5 \text{ kJ}$   $\text{H}_2 \text{O (g)}$   $\frac{1}{2} \text{ O}_2 \text{ (g)}$   $\Delta h = 71.2 \text{ kJ}$   $\text{C}_2\text{H}_6 \text{ (g)}$   $7 \text{ O}_2 \text{ (g)}$   $2 \text{ CO}_2 \text{ (g)}$   $3 \text{ H}_2 \text{O (g)}$   $\Delta h = 283 \text{ kJ}$   $\text{C}_2\text{H}_6 \text{ (g)}$

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web chemistry 120 hess s law worksheet calculate  $\Delta h$  for the reaction  $\text{C}_2\text{H}_4 \text{ (g)}$   $2 \text{ H}_2 \text{ (g)}$   $\text{C}_2\text{H}_6 \text{ (g)}$  from the following data  $\text{C}_2\text{H}_4 \text{ (g)}$   $3 \text{ O}_2 \text{ (g)}$   $2 \text{ CO}_2 \text{ (g)}$   $2 \text{ H}_2 \text{O (l)}$   $1411 \text{ kJ/mole}$   $\text{C}_2\text{H}_6 \text{ (g)}$   $7 \text{ O}_2 \text{ (g)}$   $2 \text{ CO}_2 \text{ (g)}$   $3 \text{ H}_2 \text{O (l)}$   $1560 \text{ kJ/mole}$   $\text{H}_2 \text{ (g)}$   $1/2 \text{ O}_2 \text{ (g)}$   $\text{H}_2 \text{O (l)}$   $285 \text{ kJ/mole}$

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enthalpy for naoh s hess s law says that the overall change in enthalpy for the reaction is the sum of all changes regardless of the different stages or phases of a reaction

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web thus for reaction 9 2 in which 1 mole of  $h_3o^+$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$  h 55.8 kj as shown note that the enthalpy change  $h_{products} - h_{reactants}$  is positive if heat is absorbed that is if  $h_{products}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy

exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1  $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$  2  $\text{NaOH (s)} + \text{HCl (aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{Cl}^- \text{(aq)} + \text{H}_2\text{O (l)}$  3