

Mathematically

$$\frac{P^0 - P^s}{P^0} = x_2$$

Here P^0 = Pressure of solvent in pure stage.

P^s = Pressure of solution

x_2 = Mole fraction of solute.

DERIVATION OF THE LAW

Pressure of solvent in solution

= Mole fraction of solvent \times pressure of solvent in pure stage

If solute is non-volatile

Total pressure of solution = Pressure due to solvent

= Mole fraction of solvent \times pressure of solvent in pure stage

So $P_{\text{solvent}} = x_1 \cdot P_{\text{solution}}$ Here x_1 is mole fraction of solvent.

$$\frac{P_{\text{solvent}}}{P_{\text{solution}}} = x_1$$

Subtract each term from 1 on both side.

$$1 - \frac{P_{\text{solvent}}}{P_{\text{solution}}} = 1 - x_1$$

$$\frac{P_{\text{solution}} - P_{\text{solvent}}}{P_{\text{solution}}} = x_2$$

2.11 IDEAL & NON-IDEAL SOLUTION

Sr.	Ideal solution	Non-Ideal solution
1.	A solution which obey Roult's law is called ideal solution.	A solution which do not obey Roult's law is called non-ideal solution.
2.	Interaction between components of mixture i.e.-between A&B is some as found in A-A & B-B molecule.	Interaction between components of mixture i.e.-between A&B is not some as found in A-A & B-B molecule.

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,Er. Meera Goyal,2025-02-21 1 SOLID STATE 2 SOLUTIONS 3 ELECTRO CHEMISTRY 4 CHEMICAL KINETICS 5 SURFACE
CHEMISTRY 6 GENERAL PRINCIPLES AND PROCESSES OF ISOLATION OF ELEMENTS 7 p BLOCK ELEMENTS 8 d And f
BLOCK ELEMENTS 9 COORDINATION COMPOUNDS AND ORGANOMETALLICS 10 HALOALKANES AND HALOARENES
11 ALCOHOLS PHENOLS AND ETHERS 12 ALDEHYDES KETONES AND CARBOXYLIC ACIDS 13 ORGANIC COMPOUNDS
CONTAINING NITROGEN Amines 14 BIOMOLECULES 15 POLYMERS 16 CHEMISTRY IN EVERYDAY LIFE APPENDIX 1
Important Name Reactions and Process 2 Some Important Organic Conversions 3 Some Important Distinctions Log Antilog
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VII p Block Elements Unit VIII d and f Block Elements Unit IX Coordination Compounds Unit X Haloalkanes and Haloarenes
Unit XI Alcohols Phenols and Ethers Unit XII Aldehydes Ketones and Carboxylic Acids Unit XIII Organic Compounds
Containing Nitrogen Unit XIV Biomolecules Unit XV Polymers Unit XV Polymers Unit XVI Chemistry in Everyday Life Content
1 Solid State 2 Solutions 3 Electro Chemistry 4 Chemical Kinetics 5 Surface Chemistry 6 General Principles And Processes Of
Isolation Of Elements 7 P Block Elements 8 D And F Block Elements 9 Coordination Compounds And Organometallics 10
Haloalkanes And Haloarenes 11 Alcohols Phenols And Ethers 12 Aldehydes Ketones And Carboxylic Acids 13 Organic
Compounds Containing Nitrogen 14 Biomolecules 15 Polymers 16 Chemistry In Everyday Life Appendix 1 Important Name
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 - The Rise of Digital Reading Download Surface Chemistry Class 12 Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Download Surface Chemistry Class 12 Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Download Surface Chemistry Class 12 Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Download Surface Chemistry Class 12 Notes
 - Personalized Recommendations
 - Download Surface Chemistry Class 12 Notes User Reviews and Ratings
 - Download Surface Chemistry Class 12 Notes and Bestseller Lists

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 - Participating in Virtual Book Clubs
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Download Surface Chemistry Class 12 Notes
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Download Surface Chemistry Class 12 Notes
 - Carving Out Dedicated Reading Time
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- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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