

Titration Calculation Practice Questions

1. Eggshells contain calcium carbonate. The following procedure was used to determine the percentage, by mass, of calcium carbonate in 1.13 g of eggshell.

Step 1 The 1.13 g of eggshell were crushed and added to 100.0 mL of 0.300 mol L⁻¹ HCl solution. Excess HCl remained after the reaction was complete.



Step 2 20.0 mL samples of the excess HCl were titrated with 0.10 mol L⁻¹ Na₂CO₃ solution.

- (a) Calculate the initial number of moles of HCl present before the reaction in Step 1.
(b) The equation for the titration reaction in Step 2 is shown below:



In one titration 8.35 mL of Na₂CO₃ was needed to neutralise the HCl in one 20.0 mL sample.

- (i) Calculate the number of moles of Na₂CO₃ needed to neutralise the HCl in Step 2.
(ii) Hence state the number of moles of HCl in the 20.0 mL sample titrated in Step 2.
(iii) Hence calculate the total number of moles of excess HCl that remained after the reaction with the eggshell in Step 1.
(c) Calculate the number of moles of HCl that reacted with the eggshell in Step 1.
(d) Calculate the mass of calcium carbonate in the 1.13 g of eggshell.
(e) Calculate the percentage mass of calcium carbonate in eggshell.

2. The following procedure was used to determine the concentration of hydrogen gas in a sample of air:

Step 1 1.0 × 10² L of air was bubbled through 0.0500 L of 0.300 mol L⁻¹ acidified K₂Cr₂O₇ solution. An equation for the reaction that occurred is shown below:



Excess Cr₂O₇²⁻ remained in the solution after the reaction.

Step 2 The excess Cr₂O₇²⁻ was titrated with 2.00 mol L⁻¹ Fe²⁺ solution. An equation for the reaction that occurred is shown below:



A titre value of 17.75 mL was obtained.

- (a) Calculate the number of moles of Cr₂O₇²⁻ present before the reaction with H₂ in Step 1.
(b) Calculate the number of moles of Fe²⁺ required to react with the Cr₂O₇²⁻ in Step 2.
(c) Hence calculate the number of moles of Cr₂O₇²⁻ left unreacted after Step 1.
(d) Hence calculate the number of moles of Cr₂O₇²⁻ that reacted with H₂ in Step 1.
(e) Calculate the number of moles of H₂ in the 1.0 × 10² L of air.
(f) Calculate the concentration, in %w/v, of hydrogen gas in the air.

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