

## Solutions for Gravimetric Analysis Questions

### Check for Understanding 4.1

1. Determine the solubility of AgCl using  $K_{sp}$  for AgCl and a table of initial and equilibrium concentration terms.

	$\text{AgCl}(s)$	$\rightleftharpoons$	$\text{Ag}^+(\text{aq})$	$+$	$\text{Cl}^-(\text{aq})$	
init			0.0060		0	
equil			$0.0060 + x$		$x$	where $x$ = increase in $[\text{Cl}^-]$ = solubility of AgCl

at equil,  $K_{sp} = 1.8 \times 10^{-10} = [\text{Ag}^+][\text{Cl}^-] = (0.0060 + x)(x) = 0.0060x \rightarrow x = 3.0_0 \times 10^{-8} \text{ M}$

2.  $200. \text{ mL} \times \frac{3.0_0 \times 10^{-8} \text{ mol AgCl}}{1000 \text{ mL}} \times \frac{143.32 \text{ g AgCl}}{\text{mol AgCl}} = 8.6 \times 10^{-7} \text{ g AgCl}$

### Check for Understanding 4.2

1. Often the solubility of a slightly soluble salt is increased in acidic solution because the anion combines with  $\text{H}^+$  to form a weak acid. Removal of the anion shifts the equilibrium toward more dissolution of the salt. Since Cl would form a strong acid (HCl) that is ~100% ionized in solution, it is not removed from solution. Therefore, AgCl is not more soluble in acidic solution and the presence of  $\text{HNO}_3$  does not increase its solubility.

### Check for Understanding 4.3

1. First, determine the solubility of AgCl using  $K_{sp}$  for AgCl and a table of initial and equilibrium concentration terms.

	$\text{AgCl}(s)$	$\rightleftharpoons$	$\text{Ag}^+(\text{aq})$	$+$	$\text{Cl}^-(\text{aq})$
init			0		0
equil			$x$		$x$

# Gravimetric Analysis Calculation Questions

**J Elliott**



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such as door bells and smoke detectors are shown in the figure below note explanations for common household electrical items such as three way switches and switched duplex plug outlets are below the figure notes