

CHEMISTRY MULTIPLE CHOICE QUESTIONS

A. Atoms, molecules and stoichiometry

2002 -2019



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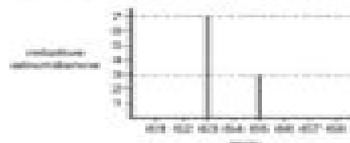
As a combustion, an alkane burns can be considered to have a fully set of 1.00 kg of methane (empirical formula CH_4).

Which mass of carbon dioxide does a person produce each day if all the carbon they take in is digested and oxidised?

- A 0.007 kg B 0.006 kg C 1.30 kg D 0.04 kg

(2002 NA/12)

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

- A 63.2 B 63.5 C 63.8 D 64.0

(2002 NA/12)

Which pairs of compounds have the same empirical formula?

- ethane and ethene
- ethane and cyclohexane
- cyclohexane and oct-1-ene

(2002 NA/12)

A stream of 10cm^3 of methane and 10cm^3 of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

Which compounds have the empirical formula CH_2O ?

- methanol
- ethanoic acid
- methyl methanoate

(2002 ON/11)

What is the number of molecules in 500 cm^3 of oxygen under room conditions?

- A 1.25×10^{23}
 B 1.25×10^{24}
 C 3.0×10^{23}
 D 3.0×10^{24}

(2004 MO/12)

In the preparation of soft margarine, glycerol trioleate



is fully hydrogenated so that, on average, one of its side-chains is converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ moiety and two side-chains are converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ moiety.

How many moles of hydrogen are required to convert one mole of glycerol trioleate into the soft margarine?

- A 4 B 5 C 6 D 8

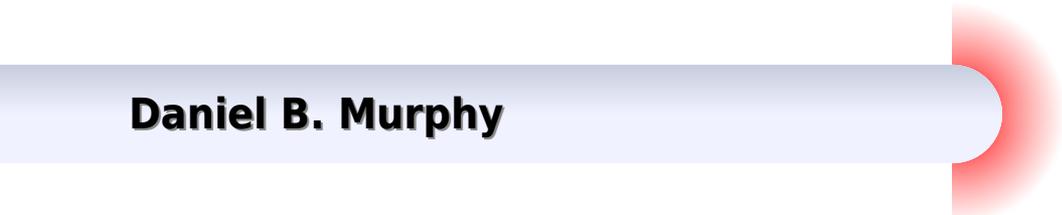
(2008 MO/12)

Use of the Data Booklet is relevant to this question.

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chemical to electrical energy applications of electrolysis reactions conductors and non conductors dry cells electrical devices circuit symbols electrolytes non electrolytes organic solvents polarization and valence electrons The chapter Electricity and Chemicals MCQs covers topics of chemical to electrical energy dry cells electrolyte non electrolyte and polarization The chapter Elements Compounds and Mixtures MCQs covers topics of elements compounds mixtures molecules atoms and symbols for elements

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Questions 226 MCQs Characteristics of Acids Bases Salts Multiple Choice Questions 115 MCQs Chemical Equilibrium
Multiple Choice Questions 46 MCQs Chemical Industries Multiple Choice Questions 67 MCQs Environmental Chemistry I
Atmosphere Multiple Choice Questions 97 MCQs Environmental Chemistry II Water Multiple Choice Questions 62 MCQs
Hydrocarbons Multiple Choice Questions 87 MCQs Organic Chemistry Multiple Choice Questions 93 MCQs Atmosphere
Multiple Choice Questions 26 MCQs The chapter Acids Bases and Salts MCQs covers topics of acids and bases concepts
Bronsted concept of acids and bases pH scale and salts The chapter Biochemistry MCQs covers topics of alcohols
carbohydrates DNA structure glucose importance of vitamin lipids maltose monosaccharide nucleic acids proteins RNA types
of vitamin vitamin and characteristics vitamin and functions vitamin and mineral vitamin deficiency vitamin facts vitamins
vitamins and supplements The chapter Characteristics of Acids Bases and Salts MCQs covers topics of concepts of acids and
bases pH measurements salts and self ionization of water pH scale The chapter Chemical Equilibrium MCQs covers topics of
dynamic equilibrium equilibrium constant and units importance of equilibrium constant law of mass action and derivation of
expression and reversible reactions The chapter Chemical Industries MCQs covers topics of basic metallurgical operations
metallurgical operations petroleum Solvay process urea and composition The chapter Environmental Chemistry I Atmosphere
MCQs covers topics of composition of atmosphere layers of atmosphere stratosphere troposphere ionosphere air pollution
environmental issues environmental pollution global warming meteorology and ozone depletion The chapter Environmental
Chemistry II Water MCQs covers topics of soft and hard water types of hardness of water water and solvent disadvantages of
hard water methods of removing hardness properties of water water pollution and waterborne diseases The chapter
Hydrocarbons MCQs covers topics of alkanes alkenes and alkynes The chapter Organic Chemistry MCQs covers topics of
organic compounds alcohols sources of organic compounds classification of organic compounds uses of organic compounds
alkane and alkyl radicals and functional groups The chapter Atmosphere MCQs covers topics of atmosphere composition air

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Table of Contents Chemistry Multiple Choice Questions With Answers

1. Understanding the eBook Chemistry Multiple Choice Questions With Answers
 - The Rise of Digital Reading Chemistry Multiple Choice Questions With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Multiple Choice Questions With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Multiple Choice Questions With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Multiple Choice Questions With Answers
 - Personalized Recommendations
 - Chemistry Multiple Choice Questions With Answers User Reviews and Ratings
 - Chemistry Multiple Choice Questions With Answers and Bestseller Lists
5. Accessing Chemistry Multiple Choice Questions With Answers Free and Paid eBooks
 - Chemistry Multiple Choice Questions With Answers Public Domain eBooks
 - Chemistry Multiple Choice Questions With Answers eBook Subscription Services

- Chemistry Multiple Choice Questions With Answers Budget-Friendly Options
- 6. Navigating Chemistry Multiple Choice Questions With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Multiple Choice Questions With Answers Compatibility with Devices
 - Chemistry Multiple Choice Questions With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Multiple Choice Questions With Answers
 - Highlighting and Note-Taking Chemistry Multiple Choice Questions With Answers
 - Interactive Elements Chemistry Multiple Choice Questions With Answers
- 8. Staying Engaged with Chemistry Multiple Choice Questions With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Multiple Choice Questions With Answers
- 9. Balancing eBooks and Physical Books Chemistry Multiple Choice Questions With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Multiple Choice Questions With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Multiple Choice Questions With Answers
 - Setting Reading Goals Chemistry Multiple Choice Questions With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Multiple Choice Questions With Answers
 - Fact-Checking eBook Content of Chemistry Multiple Choice Questions With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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