

## Chapter 4 Unit Notes

### Lesson 1: Forms of Energy

**chemical energy** stored in and released from the bonds between atoms

**electric energy** in an electric current **energy** ability to cause change **kinetic energy** due to motion

**mechanical energy** sum of potential energy and kinetic energy in a system

**nuclear energy** stored in and released from the nucleus of an atom

**potential energy** stored energy due to the interaction between objects or particles

**radiant energy** carried by electromagnetic waves

**sound energy** carried by sound waves

**speed** distance an object moves per unit of time

**thermal energy** sum of kinetic energy and potential energy of the particles that make up an object

**wave** disturbance that transfers energy from one place to another without transferring matter

#### A. Energy

1. Energy influences everything in your life.
2. Energy is the ability to cause change.

#### B. Potential Energy

1. Potential energy is stored energy due to the interaction between objects and particles.
2. Objects with potential energy have the possibility to cause change.
3. Any object has gravitational potential energy if it has mass and height above Earth's surface.
4. Chemical energy is energy that is stored in and released in the bonds between atoms.
5. Nuclear energy is the energy stored in and released from the nucleus of an atom.
  - a. The energy released from the Sun comes from nuclear fusion.
  - b. During nuclear fusion, nuclei of atoms join together and release large amounts of energy.
  - c. During nuclear fission, the nucleus of an atom breaks apart, and energy is released.
  - d. Energy used in nuclear power plants to make electricity comes from nuclear fission.

#### C. Kinetic Energy

1. Kinetic energy is energy due to motion.
2. The kinetic energy of a moving object is related to the mass and the speed of the object.
  - a. An object must have mass and speed to have kinetic energy.
  - b. If two objects have the same mass, the object that moves with greater speed has greater kinetic energy.
3. Electric energy is the energy in an electric current.

#### D. Combined Kinetic Energy and Potential Energy

1. A(n) system is a collection of interacting objects, parts, or ideas that act together as a(n) whole.
2. In science, the environment is anything that is not part of the system.
3. The sum of the potential energy and the kinetic energy in a system is mechanical energy.
4. Thermal energy is the sum of the kinetic energy and the potential energy of the particles that make up an object.

# Chapter 4 Unit Notes Lesson 1 Forms Of Energy Chemical

**Paul Franz Brandwein**



## Chapter 4 Unit Notes Lesson 1 Forms Of Energy Chemical:

*Felder's Elementary Principles of Chemical Processes, Global Edition* Richard M. Felder, Ronald W. Rousseau, Lisa G. Bullard, 2016-08-15 *Felder's Elementary Principles of Chemical Processes* prepares students to formulate and solve material and energy balances in chemical process systems and lays the foundation for subsequent courses in chemical engineering. The text provides a realistic, informative, and positive introduction to the practice of chemical engineering. This classic text has provided generations of aspiring chemical engineers with a solid foundation in the discipline: engineering problem analysis, material balances, and energy balances. Richard Felder is a recognized global leader in the field of engineering education, and this text embodies a lifetime of study and practice in effective teaching techniques. The text is in use at more than 4 out of 5 chemical engineering programs in the US.

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*Chemical Thermodynamics* Victor CM Freestone, 2025-06-13 Thermodynamics can never be made easy, but with the right approach and a consistent use of scientific terms, it can be made less opaque, and it can give a person who is prepared to try an insight into how science explains why things happen the way they do. The approach adopted in this book will give readers a better understanding of how science works together with its limitations. Unfortunately, thermodynamics, or at least some parts of it, is a subject which, apart from quantum mechanics, probably causes most confusion and bewilderment amongst scientists. The majority of students do not understand or get thermodynamics, and it is considered a hard or difficult subject. There are multiple reasons for this. There is, of course, mathematics, and many thermodynamic texts appear to be lists upon lists of differential equations. Another reason is that thermodynamics is, as often as not, poorly taught by teachers/lecturers who themselves do not understand or appreciate or have any interest in the subject. Often all three. This results not only in a lack of scientific rigor in the teaching of the subject, with the resulting confusion, and sometimes teachers/lecturers and authors just get it plain wrong; this occurs surprisingly often. However, it

need not be like this and although mathematics including calculus is required it can be kept to a relatively elementary level in order to obtain an understanding of this most important of subjects No one can pretend that the subject is easy but it can be made more accessible by a rigorous definition of terms and concepts and ensuring that a consistency of use of these definitions is maintained Highlighting the benefits of thermodynamics in practical science the text gives an intuitive grasp of the major concepts of thermodynamics such as energy and entropy Provides a new pedagogic approach to understanding and teaching chemical thermodynamics Starting with a set of basic simple assumptions about what constitutes topics such as an ideal gas theories are developed in a clear concise and accessible manner that will either answer or at the very least give an insight into a surprising range of scientific phenomena including energy heat temperature properties of gases time and quantum theory Assumes that the reader has essentially no knowledge of the subject Mathematics including calculus is kept to a relatively elementary level in order to obtain an understanding of this most important of subjects Provides the reader with a better understanding of how science works together with its limitations

**Resources in Education** ,1991 *Scott Foresman Science* ,2008 *Harcourt Science: Physical science, [grade] 5, Units E and F, teacher's ed* ,2000 *Science Makes Sense* ,1969 *Resources in Education* ,1983 *Science in Your World: Teacher edition* Jay K. Hackett,1991 *McDougal Littell Science* McDougal Littell Incorporated,2005 *The World of Matter-energy* Paul Franz Brandwein,1964 *Study Guide: Sg Concepts in Biology* Eldon D.. Enger,Enger,Frederick C.. Ross,2003 **CHEMISTRY** SILBERBERG,2003 *Psychology* James S. Nairne,2003 **Industrial & Mining Standard** ,1926 *The Mining Journal, Railway and Commercial Gazette* ,1907 **Mining Journal, Railway & Commercial Gazette** ,1907 **Scientific American** ,1909 **Basic Chemistry** Steven S. Zumdahl,1999 For a full description see catalog entry for Zumdahl *Introductory Chemistry A Foundation* 4 e **El-Hi Textbooks and Serials in Print** ,1986

## Reviewing **Chapter 4 Unit Notes Lesson 1 Forms Of Energy Chemical**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Chapter 4 Unit Notes Lesson 1 Forms Of Energy Chemical**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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