

Use your notes from the Atomic Structure program to answer the following questions.

- The atomic number tells the number of positively charged protons in the nucleus of an atom. The atom is neutral because this is also the number of negatively charged electrons in the atom.
- The mass number tells the total number of protons and neutrons in the nucleus of an atom. These particles collectively are called nucleons since both are located in the nucleus.
- Isotopes are atoms of the same element with different numbers of neutrons which results in different mass numbers.
- Write the hyphen notation of the three isotopes of hydrogen:
 - Hydrogen - 1 ${}^1_1\text{H}$
 - Hydrogen - 2 ${}^2_1\text{H}$
 - Hydrogen - 3 ${}^3_1\text{H}$
- Fill in the following chart (Hint: Isotopes may have different masses than the periodic table lists!):

Symbol	Atomic #	Mass #	# of protons	# of electrons	# of neutrons
Hg	80	201	80	80	121
Mo	42	96	42	42	54
Cu	29	64	29	29	35
C	6	14	6	6	8
	15	31	15	15	16
Pb	82	208	82	82	126
Na	11	24	11	11	13

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L Darling-Hammond



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ATOMIC STRUCTURE High School Chemistry Curriculum Bud Morgan,2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons neutrons and electrons We will also discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we will delve into the arrangement of electrons within an atom s electron cloud including energy levels sublevels and orbital shapes We will also discuss the principles of electron configuration and how it impacts an atom s chemical properties Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure We will explore periodic trends such as atomic radius ionization energy and electron affinity and discuss how these trends are related to an element s position on the table In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases We will also explore the importance of understanding atomic structure in fields such as biochemistry environmental science and materials science Throughout this book we will provide clear explanations helpful illustrations and interactive activities to reinforce key concepts Whether you are a visual learner a hands on learner or a theoretical learner this book is designed to meet your needs We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry So let s dive in and unlock the mysteries of the atom together

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