

## Standard Enthalpy of Formation ( $\Delta H^\circ_f$ ) at 298 K

Formula	KJ/mol	Formula	KJ/mol	Formula	KJ/mol
<b>Carbon</b>		<b>Chlorine</b>		<b>Sodium</b>	
C (s, graphite)	0	Cl (g)	121.3	Na (s)	0
C (s, diamond)	1.88	Cl <sub>2</sub> (g)	0	Na (g)	107.5
CH <sub>4</sub> (g)	-74.6	HCl (g)	-92.3	NaCl (s)	-411.2
CH <sub>3</sub> OH (l)	-238.6	<b>Fluorine</b>		Na <sub>2</sub> CO <sub>3</sub> (s)	-1130.7
C <sub>2</sub> H <sub>2</sub> (g)	227.4	F (g)	79.38	NaHCO <sub>3</sub> (s)	-950.8
C <sub>2</sub> H <sub>4</sub> (g)	52.4	F <sub>2</sub> (g)	0	<b>Sulfur</b>	
C <sub>2</sub> H <sub>6</sub> (g)	-84.68	HF (g)	-273.3	S <sub>8</sub> (s, rhombic)	0
C <sub>2</sub> H <sub>5</sub> OH (l)	-277.6	<b>Nitrogen</b>		S <sub>8</sub> (s, monoclinic)	0.3
C <sub>3</sub> H <sub>8</sub> (g)	-103.85	N <sub>2</sub> (g)	0	SO <sub>2</sub> (g)	-296.8
C <sub>3</sub> H <sub>6</sub> O (l, acetone)	-248.4	NH <sub>3</sub> (g)	-45.9	SO <sub>3</sub> (g)	-395.7
<b>Hydrogen</b>		NH <sub>4</sub> (s)	-365.6	H <sub>2</sub> SO <sub>4</sub> (l)	-814.0

# Standard Enthalpy Of Formation For Various Compounds

**Daniela Niemeyer**



## Standard Enthalpy Of Formation For Various Compounds:

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