

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{56}_{26}\text{Fe}$	B2																
3 (b)	$^{23}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> <th>charge</th> </tr> </thead> <tbody> <tr> <td>^{17}O</td> <td>8</td> <td>9</td> <td>10</td> <td>2-</td> </tr> <tr> <td>^{27}Al</td> <td>13</td> <td>14</td> <td>10</td> <td>3+</td> </tr> </tbody> </table>		protons	neutrons	electrons	charge	^{17}O	8	9	10	2-	^{27}Al	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
^{17}O	8	9	10	2-														
^{27}Al	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> </tr> </thead> <tbody> <tr> <td>^{32}S</td> <td>16</td> <td>16</td> <td>16</td> </tr> <tr> <td>$^{32}\text{S}^{2-}$</td> <td>16</td> <td>16</td> <td>18</td> </tr> </tbody> </table>		protons	neutrons	electrons	^{32}S	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
^{32}S	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg}(\text{s}) + \text{N}_2(\text{g}) \rightarrow \text{Mg}_3\text{N}_2(\text{s})$ <p>1 mark for correct species and state symbols 1 mark for balance</p>	B1 × 2																
6 (b)	$\text{Ca}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow \text{Ca}(\text{OH})_2(\text{aq}) + \text{H}_2(\text{g})$ <p>1 mark for correct species and state symbols 1 mark for balance</p>	B1 × 2																
6 (c)	$6\text{NaOH}(\text{aq}) + \text{Fe}_2(\text{SO}_4)_3(\text{aq}) \rightarrow 2\text{Fe}(\text{OH})_3(\text{s}) + 3\text{Na}_2\text{SO}_4(\text{aq})$ <p>1 mark for correct species and state symbols 1 mark for balance</p>	B1 × 2																

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Michael Brown



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