

# Full Stack Developer Roadmap



# Beginner Full Stack Web Development Html Css React

**Sukumar P Ganapathi Ram N**



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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{\rho} = \frac{32 \text{ g}}{1.790 \text{ g/l}} = 17.88 \text{ l}$  nonnumber

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web the molar volume of a gas from avogadro s law the volume is directly proportional to the number of moles the important point to keep in mind is that n can be moles of any gas since one mole always contain 6.022 x 10<sup>23</sup> molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6.022 x 10<sup>23</sup> particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always 6.022 x 10<sup>23</sup> but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole m<sup>3</sup> mol although it is more typical to use the units cubic decimetres per mole dm<sup>3</sup> mol for gases and cubic centimetres per mole cm<sup>3</sup> mol for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as

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web this volume is known as the molar gas volume  $V_m$  the volume of a gas may be calculated from its number of moles using volume of gas moles  $V_m$  rearranging this moles of gas frac

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web jun 8 2023  $v = 1.00 \text{ mol} \cdot 0.08206 \text{ latm mol}^{-1} \cdot 273 \text{ k} = 22.4 \text{ l}$  in other words 1 mole of a gas will occupy 22.4 l at stp

assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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web aug 3 2022 volumes of gases and liquids are measured in cubic centimetres cm<sup>3</sup> or cubic decimetres dm<sup>3</sup> or litres l 1 mole of any gas at room temperature and pressure occupies 24dm<sup>3</sup> 24 000cm<sup>3</sup> this value is known as the molar volume

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