

Chemistry 3 & 4, Chapter 4 – Volumetric Analysis TEST SOLUTIONS

Multiple Choice

Identify the choice that best completes the statement or answers the question.

Chemistry 3 & 4, Chapter 2

Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. For a substance to be suitable as a primary standard in a volumetric analysis, it should have:
- A a known formula with a low molar mass.
 - B a very low melting point
 - C no tendency to absorb moisture from the air
 - D no acid–base or redox properties.

ANSWER: C

POINTS: 0 / 1

FEEDBACK: In most volumetric analyses, the primary standard must be an acid, a base, an oxidant or a reductant. Since it must be a solid at room temperature so that it can be accurately weighed out, it needs to have a relatively high MP. Since it must be accurately weighed out, it should not absorb any substance from the air. But a low molar mass means that there will be a higher percentage of error should the measured mass differ from the true mass by, for example, 0.005 g.

REF: pp48–9

2. In a volumetric analysis of vinegar to determine the concentration of ethanoic acid present, a pair of students obtained the following results:

Initial burette reading (mL)	0.10	0.15	0.00	0.25	0.30
Final burette reading (mL)	22.70	22.60	22.50	22.73	22.60

The mean titre is:

- A 22.60 mL
- B 22.48 mL
- C 22.477 mL
- D 22.47 mL

ANSWER: B

POINTS: 0 / 1

FEEDBACK: The titres are 22.60, 22.45, 22.50, 22.48 and 22.30 mL respectively. The concordant results are 22.45, 22.48 and 22.50 mL. The mean of these is 22.48 mL, since a mean cannot have more decimal places than the results used to calculate it.

REF: pp45–6

Multiple Choice Questions Volumetric Analysis

Patrick R. Delaney, O. Ray Whittington

Multiple Choice Questions Volumetric Analysis:

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The CMA/CFM Exam Sidney J. Baxendale, Richard Coppage, Alan Levitan, 2001-05

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University Chemistry, Vol. I C. Parameswara Murthy, 2008 Cherla Parameswara Murthy Has Been Teaching At Osmania University Hyderabad For 22 Years He Is

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