

## Lab Report: Osmosis in Model Cells

4 pages | 936 words

### Abstract

The objectives of this lab were to create models of cells using dialysis tubing to demonstrate the selective permeability of the plasma membrane, study the effects of osmosis on a model cell, and predict the impact of solute concentration on osmosis. Dialysis tubing was filled with either water or varying concentrations of sucrose solutions and immersed in beakers of distilled water. The results showed that the greater the sucrose concentration inside the tubing, the greater the final mass of the system.

### Introduction

The primary goals of this lab were to explore the concept of osmosis using a model analogous to a biological cell and gain a deeper understanding of the process and nature of osmosis itself. Osmosis is a type of diffusion specific to water molecules, occurring across a selectively permeable membrane. In biological systems, the plasma membrane acts as such a barrier, allowing only certain substances like water, oxygen, and carbon dioxide to pass through while restricting the movement of large solute molecules like starch. Osmosis involves the movement of water molecules from an area of high water concentration to an area of low water concentration through a semipermeable membrane. To determine the tonicity of solutions (whether they are isotonic, hypotonic, or hypertonic compared to the cell), the total number of solute particles in the solution is considered. In this experiment, we aimed to test the hypothesis that when dialysis tubing filled with sucrose solution is immersed in a beaker of distilled water, water will move into the tubing, causing an increase in its mass. The independent variables in this experiment included the beakers of distilled water and the concentration of sucrose solutions (ranging from 0.2M to 1.0M). The dependent variables consisted of the final mass of the dialysis tubing, the change in mass of the tubing, and the percentage change in mass. The control variables included the consistent volume of sucrose solution (10ml) and distilled water used.

### Materials and Methods

#### Materials:

- Dialysis tubing
- Sucrose solutions (0.2M to 1.0M)
- Distilled water
- Beakers
- Medicine cups
- String
- Balance
- Paper towels

#### Procedure:

1. Prepare dialysis tubing for use, ensuring it is tightly sealed at one end.
2. Fill the tubing with the specified volume of sucrose solution according to the assigned concentration.
3. Securely tie off the open end of the dialysis tubing using string.
4. Rinse the filled tubing under water to check for leaks and retie if necessary.
5. Measure and record the initial mass of each sealed bag separately using a balance.
6. Place the sealed bags into corresponding beakers filled with distilled water, ensuring the bags are fully immersed.
7. Allow the bags to remain in the beakers for exactly 30 minutes to facilitate osmosis.
8. After 30 minutes, remove the bags from the beakers, gently wipe off excess liquid, and record the final mass of each bag.
9. Calculate the change in mass for each bag by subtracting the initial mass from the final mass.
10. Calculate the percentage change in mass for each bag using the formula:  $(\text{Change in Mass} / \text{Initial Mass}) \times 100$ .
11. Calculate the class average percentage change in mass for each sucrose concentration.

### Data Analysis

The experiment...

# Diffusion Osmosis Lab Report

**Scott C. Dulebohn**



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