

Worksheet 1-1 - Measuring Reaction Rates

Name - _____
Due Date - _____

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

1. A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:



A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

$$\begin{array}{r} 0.016\text{g} \\ - 0.006\text{g} \\ \hline 0.010\text{g} \end{array}$$

- a) Calculate the *Rate of Reaction* in grams of Zn consumed per second.

$$\text{rate} = \frac{\Delta g}{\Delta \text{time}} = \frac{0.010\text{g}}{20\text{sec}}$$

Answer 5.0×10^{-4}

- b) Calculate the *Rate of Reaction* in moles of Zn consumed per second.

$$5.0 \times 10^{-4} \frac{\text{g}}{\text{sec}} \times \frac{1\text{mol}}{65.4\text{g}} =$$

Answer $7.6 \times 10^{-6} \frac{\text{mol}}{\text{sec}}$

- c) Write out the complete ionic equation for the reaction.



- d) What will happen to the $[\text{H}^+]$ as the reaction proceeds?

decrease

- e) What will happen to the $[\text{Cl}^-]$ as the reaction proceeds?

nothing, it is a spectator

Rate Of Reaction Questions And Answers

RC Schank



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Table of Contents Rate Of Reaction Questions And Answers

1. Understanding the eBook Rate Of Reaction Questions And Answers
 - The Rise of Digital Reading Rate Of Reaction Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Rate Of Reaction Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Rate Of Reaction Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Rate Of Reaction Questions And Answers
 - Personalized Recommendations
 - Rate Of Reaction Questions And Answers User Reviews and Ratings

- Rate Of Reaction Questions And Answers and Bestseller Lists
- 5. Accessing Rate Of Reaction Questions And Answers Free and Paid eBooks
 - Rate Of Reaction Questions And Answers Public Domain eBooks
 - Rate Of Reaction Questions And Answers eBook Subscription Services
 - Rate Of Reaction Questions And Answers Budget-Friendly Options
- 6. Navigating Rate Of Reaction Questions And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Rate Of Reaction Questions And Answers Compatibility with Devices
 - Rate Of Reaction Questions And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Rate Of Reaction Questions And Answers
 - Highlighting and Note-Taking Rate Of Reaction Questions And Answers
 - Interactive Elements Rate Of Reaction Questions And Answers
- 8. Staying Engaged with Rate Of Reaction Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Rate Of Reaction Questions And Answers
- 9. Balancing eBooks and Physical Books Rate Of Reaction Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Rate Of Reaction Questions And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Rate Of Reaction Questions And Answers
 - Setting Reading Goals Rate Of Reaction Questions And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Rate Of Reaction Questions And Answers
 - Fact-Checking eBook Content of Rate Of Reaction Questions And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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