

One kilogram of a fluid is compressed according to the relation $(Pv = 0.25)$ where P is in bars and v in m^3/kg . The final volume is $1/4$ the initial volume. Calculate the work done on the fluid and sketch the process on a P - V diagram.

*** solution ***

∴ For closed system

$$W_{1-2} = \int P \cdot dv$$

as this fluid is compressed according to the relation $Pv = 0.25$

but $\rightarrow P$ (bars) & $v \rightarrow \frac{m^3}{kg}$
 $\therefore Pv = 0.25 \times 100 = 25 \frac{kJ}{kg}$

$\therefore P \times v = 25 \frac{kJ}{kg} \rightarrow P = \frac{25}{v} \text{ MPa}$

$$W_{1-2} = \int \frac{25}{v} \cdot dv$$

$$W_{1-2} = 25 \int \frac{dv}{v}$$

$$W_{1-2} = 25 [\ln v]_{v_1}^{v_2}$$

$$= 25 [\ln(v_2) - \ln(v_1)]$$

$$W_{1-2} = 25 \ln\left(\frac{v_2}{v_1}\right)$$

$\therefore v_2 = \frac{1}{4} v_1 \rightarrow \frac{v_2}{v_1} = \frac{1}{4}$
 $\therefore W_{1-2} = 25 \ln\left(\frac{1}{4}\right) = -34.66 \text{ kJ}$
 → which is the specific work done on the system

*** To get work

$$W = m \cdot w_{1-2} = 1 \text{ kg} \times (-34.66) \text{ kJ/kg}$$

$$\therefore W_{1-2} = -34.66 \text{ kJ}$$

(-ve) → work done on the system



2.2 A gas contained within a piston-cylinder assembly expands from an initial volume $V_1 = 0.15 \text{ m}^3$ to a final volume of $V_2 = 0.36 \text{ m}^3$ where the pressure remains constant at 0.5 MPa . Calculate the work, in kJ.

*** solution ***

For closed system shown :-

$$W_{1-2} = \int P \cdot dv$$

→ For constant Pressure Process $P_1 = P_2 = P$

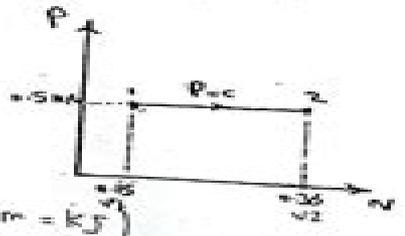
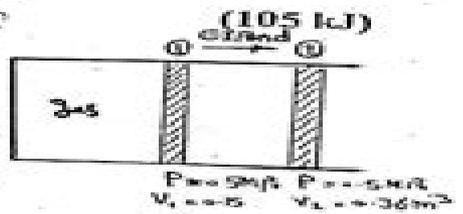
$$\therefore W_{1-2} = P \int dv = P [v]_{v_1}^{v_2}$$

$$= W_{1-2} = P (v_2 - v_1)$$

$$\therefore W_{1-2} = 0.5 \times 10^6 \text{ (N/m}^2) (0.36 - 0.15) \text{ m}^3$$

$$= 105 \text{ kJ} \quad \left(\text{kJ} \cdot \text{m}^3 = \frac{\text{kJ}}{\text{m}^3} \cdot \text{m}^3 = \text{kJ} \right)$$

(+ve) → work done by the system.



Engineering Thermodynamics Solved Problems

Mike Jess



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