



Name:

Date:

Student Exploration: Calorimetry Lab

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- The Latin word *calor* means "heat," and *meter* comes from the Greek word meaning "to measure." What do you think a **calorimeter** does?

measuring the heat of chemical reactions or physical changes as well as heat capacity

- Where have you heard the word **calorie** before? What do you think a calorie is?

calorie, a unit of energy or heat variously defined. the calorie was originally defined as the amount of heat required at a pressure of 1 standard atmosphere to raise the temperature of 1 gram of water 1 degree celsius.

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



- On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

- What was the **Final temperature** of the copper and the water?
- How much did the temperature of the copper change?
- How much did the temperature of the water change?

34.96

down 55.04

down 165.04

- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Calorimetry Lab Answers

D Keegan



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Table of Contents Calorimetry Lab Answers

1. Understanding the eBook Calorimetry Lab Answers
 - The Rise of Digital Reading Calorimetry Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Calorimetry Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Calorimetry Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Calorimetry Lab Answers
 - Personalized Recommendations
 - Calorimetry Lab Answers User Reviews and Ratings
 - Calorimetry Lab Answers and Bestseller Lists
5. Accessing Calorimetry Lab Answers Free and Paid eBooks
 - Calorimetry Lab Answers Public Domain eBooks
 - Calorimetry Lab Answers eBook Subscription Services
 - Calorimetry Lab Answers Budget-Friendly Options

6. Navigating Calorimetry Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Calorimetry Lab Answers Compatibility with Devices
 - Calorimetry Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Calorimetry Lab Answers
 - Highlighting and Note-Taking Calorimetry Lab Answers
 - Interactive Elements Calorimetry Lab Answers
8. Staying Engaged with Calorimetry Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Calorimetry Lab Answers
9. Balancing eBooks and Physical Books Calorimetry Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Calorimetry Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Calorimetry Lab Answers
 - Setting Reading Goals Calorimetry Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Calorimetry Lab Answers
 - Fact-Checking eBook Content of Calorimetry Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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