

Application of Numerical Method



- ❑ **Structural Analysis:** Structural analysis deals with a calculation of load coming on the structure and analyzing them. It is a procedure followed to find and measure the effect of external forces on physical objects like built structure elements. Numerical method is used to calculate strength, stiffness, stability and vibration of loads and beams.



Applications Of Numerical Methods In Engineering Field

Peipei Pang



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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton $h_a + b \rightleftharpoons h_b + a$ strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from ph data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral ph at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution pH is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then $\text{pH} = \log 0.100 = 1.000$ pH log 0.100 = 1.000 b titrant volume 12.50 ml

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