

# Periodic Trends Worksheet

Using the periodic table and your knowledge of periodic trends, answer the following questions:

1 Circle the best answer that has the:

- |   |                            |    |    |    |    |    |
|---|----------------------------|----|----|----|----|----|
| a | smallest electron affinity | N  | P  | As | Sb | Bi |
| b | largest electronegativity  | Ga | Ge | As | Sc |    |
| c | largest ionization energy  | O  | S  | Sc | Te | Po |
| d | lowest ionization energy   | O  | S  | Se | Te | Po |
| e | smallest atomic radius     | F  | Cl | Br | I  | At |
| f | largest atomic radius      | Li | Na | K  | Rb | Cs |

2 Low ionization energy is a characteristic of a \_\_\_\_\_

3 What does atomic radius measure? \_\_\_\_\_

4 Describe the change in the atomic size across a period and down a group.

5 Which element has the largest atomic radius? \_\_\_\_\_

6 Explain the main difference between metals and nonmetals.

7 Determine if the listed pairs of elements have similar or different properties.

- |   |                      |   |                    |
|---|----------------------|---|--------------------|
| a | Potassium & Rubidium | b | Calcium & Barium   |
| c | Sodium & Chlorine    | d | Helium & Krypton   |
| e | Phosphorus & Oxygen  | f | Lithium & Fluorine |

8 Which element has the largest electronegativity? \_\_\_\_\_

9 Which group (name) would have the lowest ionization energies?  
\_\_\_\_\_

10 Which group (name) has the highest ionization energies?  
\_\_\_\_\_

# Chemistry Practice Test Periodic Trends And Orbitals

**Léo Pomar**



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