

Answers to Chapter 3 Study Questions

1. a) $3(12.0) + 8(1.01) + 3(16.0) = 92.1 \text{ g/mole}$ b) 92.1 g c) 6.02×10^{23} molecules
 d) $0.217 \text{ moles} \times \frac{92.1 \text{ g}}{1 \text{ mole}} = 20.0 \text{ g}$
2. a) 4 atoms (one N + 3 H)
 b) $1 \text{ mole } \text{NH}_3 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 2.41 \times 10^{24} \text{ atoms}$
 c) $3.40 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 4.82 \times 10^{23} \text{ atoms}$
3. Molar mass of $\text{NaNO}_2 = 23.0 + 14.0 + 2(16.0) = 69.0 \text{ g/mole}$
 $\% \text{ Na} = 23.0/69.0 = 33.3\% \text{ Na}$; $\% \text{ N} = 14.0/69.0 = 20.3\% \text{ N}$; $\% \text{ O} = 2(16.0)/69.0 = 46.4\% \text{ O}$
33.3% Na, 20.3% N and 46.4% O.
4. a) In 100 g of this compound, there are 40.7 g C, 5.1 g H, and 54.2 g O
 $40.7 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 3.39 \text{ moles C}$ $3.39/3.39 = 1 \times 2 = 2$
 $5.1 \text{ g H} \times \frac{1 \text{ mole H}}{1.0 \text{ g H}} = 5.1 \text{ moles H}$ $5.1/3.39 = 1.5 \times 2 = 3$
 $54.2 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 3.39 \text{ moles O}$ $3.39/3.39 = 1 \times 2 = 2$
 empirical formula = **$\text{C}_2\text{H}_3\text{O}_2$**
- b) Molar mass of $\text{C}_2\text{H}_3\text{O}_2 = 2(12.0) + 3(1.0) + 2(16.0) = 59.0 \text{ g/mole}$
 $118/59.0 = 2 \rightarrow$ molecular formula = **$\text{C}_4\text{H}_6\text{O}_4$**
5. a) In 25.0 g of compound, there are 7.20 g Mg, 3.55 g C and $25.0 - (7.20 + 3.55) = 14.25 \text{ g O}$.
 $7.20 \text{ g Mg} \times \frac{1 \text{ mole Mg}}{24.3 \text{ g Mg}} = 0.296 \text{ moles Mg}$ $0.296/0.296 = 1$
 $3.55 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 0.296 \text{ moles C}$ $0.296/0.296 = 1$
 $14.25 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 0.891 \text{ moles O}$ $0.891/0.296 = 3$
 formula = **MgCO_3**
- b) $\% \text{ Mg} = 7.20/25.0 = 28.8\% \text{ Mg}$; $\% \text{ C} = 3.55/25.0 = 14.2\% \text{ C}$; $\% \text{ O} = 14.25/25.0 = 57.0\% \text{ O}$. (You should get the same result using molar mass and atomic masses.)
- c) $13.9 \text{ g compound} \times \frac{28.8 \text{ g Mg}}{100 \text{ g compound}} = 4.00 \text{ g Mg}$

Chemistry Chapter 3 Worksheet Answers

Michael Seilmaier



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Chemistry Chapter 3 Worksheet Answers Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chemistry Chapter 3 Worksheet Answers**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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web john keats 1795 1821 portrait of john keats by william hilton john keats was born in london on 31 october 1795 the eldest of thomas and frances jennings keats s four children although he died at the age of twenty five keats had perhaps the most remarkable career of any english poet