

Chemistry Notes 14.3

Three Basic Assumptions of the kinetic theory:

- 1) Gas is composed of particles, usually molecules or atoms
- 2) Particles in a gas move rapidly in constant random motion (they move in straight line paths only colliding with one another or other objects)
- 3) Collisions are perfectly elastic: kinetic energy is transferred without loss from one particle to another

Equation: $P \times V = n \times R \times T$

$R = 8.31 \text{ (L} \times \text{kPa) / (mol} \times \text{K)}$

$n = (P \times V) / (R \times T)$

Ideal Gas:

- 1) A gas is composed of particles with a mass that is low enough to be considered almost zero.
 - But a real gas has limitations to the pressure it can be compressed to before it condenses into a liquid

Real gases differ from ideal gases in the following ways:

- Real gases have small attractive and repulsive forces between particles, and ideal gases do not
- Real gas particles have a volume, and ideal gas particles do not
- Real gas particles collide inelastically and lose energy with collisions, and ideal gas particles collide elastically
- Real gases are not infinitely compressible

Ideal gases don't exist because they do take up space, and there are attractive forces

Diffusion: molecules moving from areas of high concentration to low concentration

- Perfume

Effusion: Gas escaping through tiny hole in container

- Both of these depend on the molar mass of the particle, which determines the speed
- The rate of effusion and diffusion is inversely proportional to the square root of the molar mass

To solve Graham's law problems, put whatever you're looking for in the numerator
6777

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Marcel A. Müller



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Table of Contents Chemistry Ch 14 Guided Reading

1. Understanding the eBook Chemistry Ch 14 Guided Reading
 - The Rise of Digital Reading Chemistry Ch 14 Guided Reading
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Ch 14 Guided Reading
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Ch 14 Guided Reading
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Ch 14 Guided Reading
 - Personalized Recommendations
 - Chemistry Ch 14 Guided Reading User Reviews and Ratings
 - Chemistry Ch 14 Guided Reading and Bestseller Lists
5. Accessing Chemistry Ch 14 Guided Reading Free and Paid eBooks
 - Chemistry Ch 14 Guided Reading Public Domain eBooks

- Chemistry Ch 14 Guided Reading eBook Subscription Services
- Chemistry Ch 14 Guided Reading Budget-Friendly Options
- 6. Navigating Chemistry Ch 14 Guided Reading eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Ch 14 Guided Reading Compatibility with Devices
 - Chemistry Ch 14 Guided Reading Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Ch 14 Guided Reading
 - Highlighting and Note-Taking Chemistry Ch 14 Guided Reading
 - Interactive Elements Chemistry Ch 14 Guided Reading
- 8. Staying Engaged with Chemistry Ch 14 Guided Reading
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Ch 14 Guided Reading
- 9. Balancing eBooks and Physical Books Chemistry Ch 14 Guided Reading
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Ch 14 Guided Reading
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Ch 14 Guided Reading
 - Setting Reading Goals Chemistry Ch 14 Guided Reading
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Ch 14 Guided Reading
 - Fact-Checking eBook Content of Chemistry Ch 14 Guided Reading
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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