

Question number	Answer	Marks	Guidance
1 (a)	$n(\text{NaOH}) = 0.125 \times 22.40/1000 = 2.80 \times 10^{-3} \text{ mol}$	B1	
1 (b)	$n(\text{NaOH}) = 0.5 \times 2.80 \times 10^{-3} = 1.40 \times 10^{-3} \text{ mol}$	B1	
1 (c)	$c = 1.40 \times 10^{-3} \times 1000/25.0 = 0.0560 \text{ mol dm}^{-3}$	B1	
2	N in $\text{NH}_3$ has been oxidised from $-3$ to $0$ in $\text{N}_2$ Cu in $\text{CuO}$ has been reduced from $+2$ to $0$ in $\text{Cu}$	B1 B1	
3 (a) (i)	Mg has lost 2 electrons and has been oxidized Fe has lost 3 electrons and has been reduced	B1 B1	
3 (a) (ii)	iron(III) nitrate(V)	B1	
3 (b) (i)	Mn has been reduced from $+4$ in $\text{MnO}_2$ to $+2$ in $\text{MnCl}_2$	B2	
3 (b) (ii)	In $\text{HCl}$ , Cl has oxidation number of $-1$ . Cl has been oxidised from $-1$ in $\text{HCl}$ to $0$ in $\text{Cl}_2$ Cl in $\text{HCl}$ has an oxidation number of $-1$ which is unchanged in $\text{MnCl}_2$	B1 B1	
4 (a) (i)	For each equation, 1 mark for balanced equation, 1 mark for state symbols. $\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ $\text{CaO}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l})$ $\text{Ca}(\text{OH})_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$	B1 x 2 B1 x 2 B1 x 2	
4 (a) (ii)	An acid has been neutralised by a base to form water	B1	
4 (a) (iii)	$\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$	B1	
4 (b) (i)	Reduction is increase in oxidation number and gain of electrons Oxidation is decrease in oxidation number and loss of electrons	B1 B1	
4 (b) (ii)	$\text{Ca}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{g}) + \text{H}_2(\text{g})$	B2	
4 (b) (iii)	Ca has been oxidised from $0$ to $+2$ H has been reduced from $1+$ to $0$ .	B1 B1	
5 (a)	$\text{WO}_3 + 3\text{H}_2 \rightarrow \text{W} + 3\text{H}_2\text{O}$	B1	

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## **Table of Contents Access Chapter 4 Answers**

1. Understanding the eBook Access Chapter 4 Answers
  - The Rise of Digital Reading Access Chapter 4 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Access Chapter 4 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Access Chapter 4 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Access Chapter 4 Answers

- Personalized Recommendations
  - Access Chapter 4 Answers User Reviews and Ratings
  - Access Chapter 4 Answers and Bestseller Lists
5. Accessing Access Chapter 4 Answers Free and Paid eBooks
    - Access Chapter 4 Answers Public Domain eBooks
    - Access Chapter 4 Answers eBook Subscription Services
    - Access Chapter 4 Answers Budget-Friendly Options
  6. Navigating Access Chapter 4 Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Access Chapter 4 Answers Compatibility with Devices
    - Access Chapter 4 Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Access Chapter 4 Answers
    - Highlighting and Note-Taking Access Chapter 4 Answers
    - Interactive Elements Access Chapter 4 Answers
  8. Staying Engaged with Access Chapter 4 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Access Chapter 4 Answers
  9. Balancing eBooks and Physical Books Access Chapter 4 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Access Chapter 4 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Access Chapter 4 Answers
    - Setting Reading Goals Access Chapter 4 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Access Chapter 4 Answers

- Fact-Checking eBook Content of Access Chapter 4 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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