

## Jharkhand University of Technology, Ranchi

1st Semester Diploma Examination, 2022

Subject : Communication Skill-I

Time Allowed : 3 Hours

Subject Code : 101

Full Marks : 80

Pass Marks : 26

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*Answer in your own words.**Answer all questions.**Read instruction carefully.*

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1. Answer any four from the following:

2x4=8

- (i) What is important to become a good public speaker?
- (ii) How did Shaw come out of his nervousness?
- (iii) According to Dr. Kalam what is essential for happiness, peace and prosperity?
- (iv) State the qualities of a magnetic leader.
- (v) Why was Gandhiji against modern machinery?
- (vi) How was the Singer Sewing Machine invented?
- (vii) Who inspired Kalpana Chawla?
- (viii) What is the technical definition of fossil fuel? What is its importance?
- (ix) Which areas did Kiran Bedi prefer to work?
- (x) How does A. G. Gardiner elaborate the beauty of sunrise and sunset?

2. Write a short paragraph on any one in about 200 words:

1x8=8

- (i) Fit India Movement
- (ii) Swachh Bharat Abhiyan
- (iii) A burning candle can light another
- (iv) If you look to do something it is always possible

# Diploma Second Semester Mathematics Questions Paper

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## **Diploma Second Semester Mathematics Questions Paper:**

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web 1 the chemical equations in model 1 contain the phase notations s l g and aq write each symbol by each meaning dissolved in water liquid solid gas aq l s g 2 based on the

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web sep 16 2022 classify a chemical reaction as a synthesis decomposition single replacement double replacement or a combustion reaction predict the products of simple reactions the chemical reactions we have described are only a tiny sampling of the infinite number of chemical reactions possible

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web identifying chemical reaction types key part 1 identify the type of reaction 1  $\text{Na}_3\text{PO}_4 + 3\text{KOH} \rightarrow 3\text{NaOH} + \text{K}_3\text{PO}_4$  double displacement 2  $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$  double displacement 3  $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$  combustion 4  $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$  single displacement 5  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$  decomposition 6  $\text{P}_4 + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$

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web 1  $3\text{NaBr} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{HBr}$  type of reaction double displacement 2  $3\text{Ca(OH)}_2 + 2\text{Al}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{Al(OH)}_3$  type of reaction double displacement 3  $3\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow 2\text{Fe} + 3\text{MgO}$  type of reaction single displacement 4  $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$  type of reaction combustion 5  $2\text{PbSO}_4 + 2\text{PbSO}_3 + \text{O}_2$  type of reaction

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