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Karen Morrison & Nick Hamshaw



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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + 5/2 \text{ O}_2 \text{ (g)}$ $\Delta h = 2 \text{ H}_2 \text{ (g)} + \text{C}_2\text{H}_6 \text{ (g)}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} \rightarrow 1/2 \text{ O}_2 \text{ (g)}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)} + 7/2 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (g)}$ $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 2 \text{ H}_2\text{O (l)}$ $\Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)} + 7/2 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (l)}$ $\Delta h = 1560 \text{ kJ}$
 $\text{H}_2 \text{ (g)} + 1/2 \text{ O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$ $\Delta h = 285 \text{ kJ}$
 mole

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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2 which acid is produced in the stomach a HNO_3 b CH_3COOH c

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web nov 13 2022 suggest simple tests you could carry out to determine if an unknown substance is an acid or a base state the chemical definitions of an acid and a base in terms of their behavior in water write the formula of the salt formed when a given acid and base are combined

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web the ionization of weak acids and weak bases many acids and bases are weak that is they do not ionize fully in aqueous solution a solution of a weak acid in water is a mixture of the nonionized acid hydronium ion and the conjugate base of the acid with the nonionized acid present in the greatest concentration

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web oct 10 2022 correct answer d salt water explanation neutralization is a chemical reaction between an acid and a base resulting in the formation of a salt and water the product of neutralization is therefore salt water as the acid and base combine to form a salt compound and water molecules rate this question 4