

K_w CALCULATION

$$K_w = [H^+][OH^-]$$



$$\Rightarrow [H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{0.46}$$

$$pH = -\log[H^+]$$

$$= -\log(2.17391 \times 10^{-14})$$
$$= 13.66$$

pH to 2 d.p.!

A-LEVEL CHEMISTRY

Calculations For A Level Chemistry

Michael Hoelscher



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Table of Contents Calculations For A Level Chemistry

1. Understanding the eBook Calculations For A Level Chemistry
 - The Rise of Digital Reading Calculations For A Level Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Calculations For A Level Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Calculations For A Level Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Calculations For A Level Chemistry
 - Personalized Recommendations
 - Calculations For A Level Chemistry User Reviews and Ratings
 - Calculations For A Level Chemistry and Bestseller Lists

5. Accessing Calculations For A Level Chemistry Free and Paid eBooks
 - Calculations For A Level Chemistry Public Domain eBooks
 - Calculations For A Level Chemistry eBook Subscription Services
 - Calculations For A Level Chemistry Budget-Friendly Options
6. Navigating Calculations For A Level Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Calculations For A Level Chemistry Compatibility with Devices
 - Calculations For A Level Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Calculations For A Level Chemistry
 - Highlighting and Note-Taking Calculations For A Level Chemistry
 - Interactive Elements Calculations For A Level Chemistry
8. Staying Engaged with Calculations For A Level Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Calculations For A Level Chemistry
9. Balancing eBooks and Physical Books Calculations For A Level Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Calculations For A Level Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Calculations For A Level Chemistry
 - Setting Reading Goals Calculations For A Level Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Calculations For A Level Chemistry
 - Fact-Checking eBook Content of Calculations For A Level Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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