

## Periodic Trends Lab

Periodic law states that when the elements are arranged in order of increasing atomic number, there is a periodic repetition of their chemical and physical properties. Elements within the same group have the most similar chemical and physical properties, although not identical. Elements in neighboring groups may also have similar chemical and physical properties, but to a lesser degree. Trends in properties occur within groups and across periods.

**PURPOSE:** In this activity, you will

- Interpret group and periodic trends in atomic radius, electronegativity, and ionization energy.
- Graph these trends.

**PROCEDURE:**

1. Complete the pre-lab section.
2. Using the data from **Table S** of your reference tables, prepare the graphs listed below. Staple your graphs to the back of this lab packet.

*On each graph do your best:*

- USE PENCIL!
- Put an appropriate title and labels for each axis.
- Omit the transition elements.
- Connect the points you plot.
- Connect the points representing Group 1 elements with a green line.
- Connect the points representing Group 2 elements with a blue line.
- Connect the points representing Group 17 elements with a red line.
- Draw a triangle around the points representing the elements of Period 3.

GRAPH 1. Atomic radius vs. atomic number for elements 1-54.

GRAPH 2. Electronegativity vs. atomic number for elements 1-54.

GRAPH 3. Ionization energy vs. atomic number for elements 1-54.

3. Complete the conclusion section.

# Periodic Trends Lab Question Answers

**Patrick Vollmar**



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