

7-2 Review and Reinforcement

Covalent Bonding

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided.

- | | |
|---|---|
| ___ covalent
___ true
___ 4 shared (2 pairs)
___ true
___ structural
___ lone/non-bonding
___ equally
___ triple | 1. A group of atoms united by <u>ionic</u> bonds is called a molecule.
2. A covalent bond is formed by a <u>shared</u> pair of electrons.
3. A double covalent bond consists of <u>two</u> shared electrons.
4. A <u>molecular</u> formula tells you how many atoms are in a single molecule of the compound.
5. The <u>empirical</u> formula for a molecule specifies which atoms are bonded to each other in the molecule.
6. A pair of electrons not involved in bonding is called a <u>shared electron pair</u> .
7. In nonpolar covalent bonds, the electrons are shared <u>unequally</u> between two atoms.
8. Lewis structures use a triple dash to represent a <u>double</u> bond. |
|---|---|

Use the information in Figure 7-19 of the textbook to calculate the electronegativity differences between each pair of elements listed in the chart below. Then indicate whether the bond would be nonpolar covalent, polar covalent, or ionic. Write your answers in the chart.

		Electronegativity Difference	Type of Bond
9.	H—O	1.4	polar covalent
10.	C—H	.4	non-polar covalent
11.	K—F	3.2	ionic
12.	N—H	.9	polar covalent
13.	Na—F	3.1	ionic
14.	O—Cl	.5	polar covalent

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