

3. Atoms, elements and compounds



Atomic mass = 23

Proton number = 11

Neutron number = 12 (23 - 11)

mass of electron = $\frac{1}{1836}$

ISOTOPE: atoms of the same element with the same proton number but a different neutron number

- Have the same properties due to the same number of electrons on the outer shell

Radioisotope

- testing for leaks in pipes
- treating cancer

IONIC BONDING

Bonding: oppositely charged ions are strongly attracted by electrostatic force
+ involves electron transfer

metal + non metal

- metals give electrons to non-metals

Properties: strong bonds due to attractive forces

- conduct electricity
- high mp/bp
- charged particles
- soluble in water

Macromolecular compounds

SILICON (IV) OXIDE

Properties



COVALENT BONDING

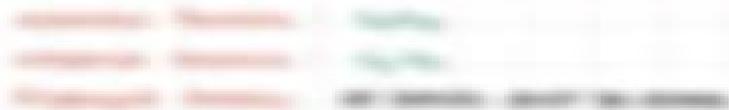
Bonding: nuclei of atoms and the shared pair of electrons are strongly attracted by electrostatic forces

non metal + non metal

Properties: weak bonds = low mp/bp

- insoluble
- don't conduct

3. Bonding



... the strong force of attraction between the ...
... the ...
... the ...

A large box containing a list of properties and examples for different types of bonding. It lists properties like melting point, boiling point, and conductivity for ionic, covalent, and metallic bonding, along with examples like NaCl, diamond, and copper.

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