

KEY

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of Ka for the acid HBrO. (4 marks) $pH = 4.65$ so $[H_3O^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} M$



| | | | | |
|-----|-------------------------------|--|-------------------------|-------------------------|
| [I] | 0.25 | | 0 | 0 |
| [C] | -2.239×10^{-5} | | $+2.239 \times 10^{-5}$ | $+2.239 \times 10^{-5}$ |
| [E] | $0.25 - 2.239 \times 10^{-5}$ | | 2.239×10^{-5} | 2.239×10^{-5} |

4

$$K_a = \frac{[H_3O^+][BrO^-]}{[HBrO]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) $pH = 2.355$ $[H_3O^+] = \text{antilog}(-2.355) = 0.004416 M$



| | | | | |
|-----|------------------|--|-------------|-------------|
| [I] | C_0 | | 0 | 0 |
| [C] | -0.004416 | | $+0.004416$ | $+0.004416$ |
| [E] | $C_0 - 0.004416$ | | 0.004416 | 0.004416 |

4

$$K_a = \frac{[H_3O^+][C_6H_5COO^-]}{[C_6H_5COOH]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 M$$

Answer $C_0 = 0.30 M$

5. Find the value of Kb for the oxalate ion ($C_2O_4^{2-}$). (1 mark)

$$K_b(C_2O_4^{2-}) = \frac{K_w}{K_a(HC_2O_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

1

Answer $K_b = 1.6 \times 10^{-10}$

6. The value of Kb for the weak base methylamine (CH_3NH_2) is 4.4×10^{-4} . Calculate the value of Ka for the acid $CH_3NH_3^+$. (1 mark)

$$K_a(CH_3NH_3^+) = \frac{K_w}{K_b(CH_3NH_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

1

Answer $K_a = 2.3 \times 10^{-11}$

10

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