

Introduction

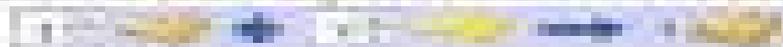
This course is a general chemistry course for students who are interested in pursuing a degree in a science or engineering field. The course covers the basic principles of chemistry, including atomic structure, chemical bonding, and chemical reactions. The course is designed to provide students with a solid foundation in chemistry, which is essential for many careers in science and technology.

Prerequisites: Students must have completed high school chemistry or a college-level chemistry course with a grade of C or better. Students who do not meet these prerequisites should consult with an advisor for more information.

Part 1: Atomic Structure



1. Define the terms atom, element, and compound. Explain the difference between an atom and a molecule. Describe the structure of an atom, including the nucleus and the electron cloud.
2. Explain the periodic table of elements. Describe the trends in atomic size, ionization energy, and electronegativity.
3. Calculate the atomic weight of an element from its isotopic composition.



4. Calculate the number of protons, neutrons, and electrons in an atom of a given element.
5. Calculate the number of protons, neutrons, and electrons in an ion of a given element.

Element	Atomic Weight	Atomic Number	Number of Protons	Number of Neutrons	Number of Electrons
H	1.008	1	1	1	1
He	4.003	2	2	2	2
Li	6.941	3	3	3	3
Be	9.012	4	4	4	4

6. Calculate the number of protons, neutrons, and electrons in an ion of a given element. The following table shows the number of protons, neutrons, and electrons in some common ions.

Na ⁺	11	11	10
Cl ⁻	17	17	18

Part 2: Chemical Bonding



7. Explain the difference between ionic and covalent bonding. Describe the formation of an ionic bond and a covalent bond.
8. Calculate the lattice energy of an ionic compound from its thermodynamic data.



9. Calculate the number of protons, neutrons, and electrons in an ion of a given element. The following table shows the number of protons, neutrons, and electrons in some common ions.

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