

GOUACHE Painting Tutorial

Let's paint
simple
GRASS!

by Hermina Utami



light tone



mid tone



dark tone

These colors are a mixture of light yellow + deep green.



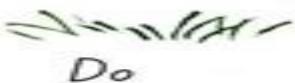
2 Add the second color (mid tone).*



3 Add the third color (dark tone).*

*when the previous layer is completely dry OR still wet

4 Use the same color as the base layer to paint the grass.



Do



Don't



5 Then, decorate it with flowers —white or any other colors.

Painting Techniques Manual Reference

Mark Steyvers



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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the h^+ ions from the acid and the oh^- ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web other articles where exothermic reaction is discussed alkali metal reactions with nonmetals the reactions are highly exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 via and via of the periodic table sulfides can be formed by the direct reaction of the alkali metals with

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm³ 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o l}$ for which $\Delta h = 57 \text{ kJ mol}^{-1}$ eg 2 $\text{h}_2\text{so}_4 \text{ aq} + \text{koh aq} \rightarrow \text{ks}_2\text{o}_7 \text{ aq} + \text{h}_2\text{o l}$ for which $\Delta h = 57 \text{ kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o aq}$ you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts

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web 1 label the pyramid levels in model 1 with the following primary producers primary consumers secondary consumers and tertiary consumers 2 the arrows in model 1

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pyramid of energy is the only type of ecological pyramid which is always upright as the energy flow in a food chain is always unidirectional also with every increasing trophic level some energy is lost into the see more

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web refer to model 1 a how much energy per year do the caterpillars in model i obtain from eating the leaves in a square meter of the oak tree y ooo tcti b what percentage of

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in this particular type of ecological pyramid each level takes into account the amount of biomass produced by each trophic

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