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web the enol keto equilibrium constant  $k_{eq}$  enol keto and enol are functions of temperature their values can be determined by integrating the enol and keto signal areas table 1 summarizes the equilibrium constant and percent enol concentration values calculated for the different combinations of enol to keto ratio note that

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web the relative amounts of products and reactants at equilibrium is characterized by the equilibrium constant for this reaction  $K_{eq}$  is defined as  $\frac{[enol]_{eq}}{[keto]_{eq}}$  where the subscript eq indicates that all concentrations are measured at equilibrium

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web because the system exists essentially completely as the thiol isomer a carbon acid acidity constant for ionization starting with the thio keto form as the initial state  $K_j$  could not be measured and a keto enol equilibrium constant  $K_{te}$  could not be determined

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web keto enol tautomerism because of the acidity of  $\alpha$  hydrogens carbonyls undergo keto enol tautomerism tautomers are rapidly interconverted constitutional isomers usually distinguished by a different bonding location for a labile hydrogen atom and a differently located double bond

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web jan 27 2016 several factors affect the position of keto enol equilibrium defined here as  $K_{e/k}$  enol keto including structure steric bulk conjugation electron withdrawing donating groups resonance temperature and solvent a judiciously selected set of compounds that have a common 1 3 dicarbonyl moiety with progressively

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web experimental methods detailed experimental procedure used for this experiment is given in the chem 417 laboratory manual the experiment 1 nmr determination of keto enol equilibrium constant each student prepared 4 tubes with sample solutions and two nmr tubes with solvent only one with deuterated chloroform and one with the methanol

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web abstract the keto enol equilibrium constants of acetylacetone ethyl acetoacetate and ethyl benzoylacetate in water at 25  $^{\circ}C$  are determined by studying the influence of surfactants on their uv vis spectra following the method applied to benzoylacetone published recently

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web nov 1 2007 the use of proton nmr to determine the equilibrium position of tautomeric 1 3 dicarbonyl compounds in

various solvents has been a classic physical chemistry experiment we are presenting an expansion of the excellent description of this experiment by garland shoemaker and nibler

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web using nmr spectroscopy the equilibrium values for different ketones which undergo tautomerization were found by determining relative peak values for each species the value for acetylacetone was found to be  $k_c = 6.622$  and for ethyl acetoacetate was  $k_c = 7.510$

**solved nmr determination of keto enol equilibrium constants** - Jul 19 2022

web nmr determination of keto enol equilibrium constants the nmr spectrum below is for a acetylacetone with  $CDCl_3$  as solvent the photos depict the complete nmr spectrum as well more zoomed in areas of each peak demonstrate how to calculate the equilibrium constant for keto to enol conversion  $k_c$