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Determining Molecular Formulas (True Formulas)

1. The empirical formula of a compound is  $\text{NO}_2$ . Its molecular mass is 92 g/mol. What is its molecular formula?

$$\begin{array}{r} \text{NO}_2 \\ 16.00\text{g} \\ 16.00\text{g} \\ + 14.00\text{g} \\ \hline 46.00\text{g} \end{array} \quad \frac{92\text{g}}{46.0\text{g}} = 2 \quad \text{N}_2\text{O}_4$$

2. The empirical formula of a compound is  $\text{CH}_2$ . Its molecular mass is 70 g/mol. What is its molecular formula?

$$\begin{array}{r} \text{CH}_2 \\ 12.01\text{g} \\ 1.01\text{g} \\ + 1.01\text{g} \\ \hline 14.03\text{g} \end{array} \quad \frac{70\text{g}}{14.03\text{g}} = 4.98 = 5 \quad \text{C}_5\text{H}_{10}$$

3. A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.3% oxygen. Its molecular mass is 60. g/mol. What is its molecular formula?

$$\begin{array}{l} 40.0\% \text{ C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 3.33\text{mol C} \frac{3.33}{3.33} = 1 \\ 6.7\% \text{ H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 6.63\text{mol H} \frac{6.63}{3.33} = 2 \\ 53.3\% \text{ O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 3.34\text{mol O} \frac{3.34}{3.33} = 1 \end{array} \quad \begin{array}{l} \text{CH}_2\text{O} = 50.02\text{g} \\ \frac{60\text{g}}{30\text{g}} = 2 \\ \text{C}_2\text{H}_4\text{O}_2 \end{array}$$

4. A compound is 64.9% carbon, 13.5% hydrogen and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

$$\begin{array}{l} 64.9\% \text{ C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 5.40\text{mol C} \frac{5.40}{1.35} = 4 \\ 21.6\% \text{ O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 1.35\text{mol O} \frac{1.35}{1.35} = 1 \\ 13.5\% \text{ H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 13.4\text{mol H} \frac{13.4}{1.35} = 10 \end{array} \quad \begin{array}{l} \text{C}_4\text{H}_{10}\text{O} \\ \uparrow \\ \text{empirical} \\ \uparrow \\ \text{molecular} \end{array}$$

5. A compound is 54.5% carbon, 9.1% hydrogen and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

$$\begin{array}{l} 54.5\% \text{ C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 4.53\text{mol C} \frac{4.53}{2.28} = 2 \\ 9.1\% \text{ H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 9.00\text{mol H} \frac{9.00}{2.28} = 4 \\ 36.4\% \text{ O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 2.28\text{mol O} \frac{2.28}{2.28} = 1 \end{array} \quad \begin{array}{l} \text{C}_2\text{H}_4\text{O} = 44 \\ \frac{88}{2} = 44 \\ \text{C}_4\text{H}_8\text{O}_2 \end{array}$$

# Molecular Formula Worksheet Answer Key

**Camilla Rothe**



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